#### EMPIRICAL AND MOLECULAR FORMULAE

# EMRIRICAL FORMULA - This represents the simplest whole number ratio of atoms within a molecule

# MOLECULAR FORMULA - This represents the exact numbers of each atom in a molecule

Empirical formulae can be calculated from either % mass composition data or mass composition data. Molecular formula can then be calculated if the molecular mass of the compound is known.

#### EXAMPLE 1 - from % composition

A substance containing carbon and hydrogen and oxygen only contains 73.47% carbon and 10.20% hydrogen by mass, its relative molecular mass is 98. Calculate its molecular formula.

Step 1 calculate empirical formula

ATOMS	CARBON	HYDROGEN	OXYGEN
PERCENTAGE MASS	73.47	10.20	(100 - (73.47 + 10.20) = 16.33
ATOMIC MASS	12	1	16
MASS/ATOMIC MASS	6.1225	10.20	1.020
RATIO(divide by smallest)	6	10	1

Empirical formula = C.H.O

Step 2 use Mo to find molecular formula

 $M_{\rm P}$  = mass ( $C_6H_{10}O$ )  $\times$  n

98 = (98) x n n = 1

molecular formula =

CHO

#### EXAMPLE 2 - from mass data

A substance contains 12.28g nitrogen, 3.51g hydrogen, 28.07g sulphur and 56.10g oxygen, its Me is 228 .Calculate its molecular formula.

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**E Durkheim** 

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boy more sauce servus cui salvius hoc imperavit statim exiit brevi regressus garum veneno mixtum intulit atque in belimici pateram effudit the slave to whom salvius had ordered left at once having returned in a short time he brought in the sauce mixed with poison and poured it out onto belimicus s bowl

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