

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K? $PV = nRT$
 $V = 120 \text{ L}$
 $P = 2.3 \text{ atm}$
 $T = 340 \text{ K}$
 $n = ?$

$$(2.3 \text{ atm})(120 \text{ L}) = n (0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(340 \text{ K})$$

$$\boxed{n = 9.9 \text{ mol}}$$
- 2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200° C, what is the pressure inside the container?
 $V = 50. \text{ L}$
 $n = 45 \text{ mol}$
 $T = 200.^{\circ}\text{C} = 473 \text{ K}$
 $P = ?$

$$(P)(50. \text{ L}) = (45 \text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(473 \text{ K})$$

$$\boxed{P = 30 \text{ atm}} \quad \text{or} \quad \boxed{4000 \text{ kPa}}$$
- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
 $V = 1.0 \text{ L}$
 $n = 2 \text{ mol}$
 $T = 1400.^{\circ}\text{C} = 1673 \text{ K}$
 $P = ?$

$$(P)(1.0 \text{ L}) = (2 \text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(1673 \text{ K})$$

$$\boxed{P = 275 \text{ atm} = 280 \text{ atm}}$$

$$\text{or } 28,000 \text{ kPa}$$
- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?
 $V = 30. \text{ L}$
 $T = 300. \text{ K}$
 $P = 200. \text{ atm}$
 $n = ?$

$$(200. \text{ atm})(30. \text{ L}) = n (0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(300. \text{ K})$$

$$\boxed{n = 243.6 \text{ mol} = 240 \text{ mol}}$$
- 5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 3.0 moles of oxygen gas at a pressure of 1.0 atmosphere, what is the temperature of the balloon?
 $V = 100. \text{ L}$
 $n = 3 \text{ mol}$
 $P = 1 \text{ atm}$
 $T = ?$

$$(1 \text{ atm})(100. \text{ L}) = (3 \text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(T)$$

$$\boxed{T = 406 \text{ K} = 133.^{\circ}\text{C} = 130.^{\circ}\text{C}}$$

Ideal Gas Law Practice Work Sheet

A Gutmann



Ideal Gas Law Practice Work Sheet:

Holt Chemistry Ralph Thomas Myers, 2004 *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **A Guide to Teaching in the Active Learning Classroom** Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03

While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment for administrators planning to create ALCs or experimenting with provisionally designed rooms and for faculty developers helping teachers transition to using these new spaces.

American Journal of Physics, 2005 *Spreadsheets for Chemists* Gordon Filby, 1995

A practical guide. *Spreadsheets for Chemists* shows chemists of all levels how to use spreadsheet programs in their daily work. It highlights the possibilities provided by Lotus 1 2 3, the most widely used spreadsheet program in the sciences. Apart from hundreds of example fragments, it features a detailed discussion of the most relevant functions and all the macro commands. An accompanying diskette containing 57 worksheets involving many different fields of chemical research and teaching. An extensive glossary of spreadsheet terms. Three appendices covering 1 2 3's competitors and add-in packages, the use of Windows-based spreadsheets, and how what-if analysis and back-solving is applied. Although the disk examples were developed for Lotus 1 2 3 DOS Versions 2 x x 2 4, the worksheets are compatible with the newer Windows versions and those of their main competitors, Borland's Quattro Pro and Microsoft's Excel. Several compatible spreadsheets, *AsEasyAs Pro* might also be used as inexpensive alternatives. The author is a senior scientist at the Nuclear Research Centre in Karlsruhe, Germany. He has been using spreadsheet software for nearly ten years successfully in a variety of chemical problems. Simulations and Student

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Table of Contents Ideal Gas Law Practice Work Sheet

1. Understanding the eBook Ideal Gas Law Practice Work Sheet
 - The Rise of Digital Reading Ideal Gas Law Practice Work Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Ideal Gas Law Practice Work Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ideal Gas Law Practice Work Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal Gas Law Practice Work Sheet
 - Personalized Recommendations
 - Ideal Gas Law Practice Work Sheet User Reviews and Ratings
 - Ideal Gas Law Practice Work Sheet and Bestseller Lists
5. Accessing Ideal Gas Law Practice Work Sheet Free and Paid eBooks
 - Ideal Gas Law Practice Work Sheet Public Domain eBooks

- Ideal Gas Law Practice Work Sheet eBook Subscription Services
- Ideal Gas Law Practice Work Sheet Budget-Friendly Options
- 6. Navigating Ideal Gas Law Practice Work Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Ideal Gas Law Practice Work Sheet Compatibility with Devices
 - Ideal Gas Law Practice Work Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ideal Gas Law Practice Work Sheet
 - Highlighting and Note-Taking Ideal Gas Law Practice Work Sheet
 - Interactive Elements Ideal Gas Law Practice Work Sheet
- 8. Staying Engaged with Ideal Gas Law Practice Work Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ideal Gas Law Practice Work Sheet
- 9. Balancing eBooks and Physical Books Ideal Gas Law Practice Work Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ideal Gas Law Practice Work Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ideal Gas Law Practice Work Sheet
 - Setting Reading Goals Ideal Gas Law Practice Work Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ideal Gas Law Practice Work Sheet
 - Fact-Checking eBook Content of Ideal Gas Law Practice Work Sheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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