POGIL: Períodic Table Trends

(Adapted from Rush Henrietta CSD)

Wilhard

The Periodic Table is one of the greatest inventions in the history of man. It allows scientists to predict physical and chemical properties of the elements. Dimitri Mendeleev (a Russian scientist) and Robert Mosley (a British chemist) put together this table in the late 1800's based on properties of the elements known at the time. The trends in these properties as you go across periods and down groups is the subject of this Chem POGIL.

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Success Criteria:

- Understand the meaning of
 - atomic radius, reactivity, electronegativity, ionization energy
- Recognize trends in atomic radius, reactivity, ionization energy and electronegativity as you go across periods and down groups.

Resources:

Reference Table S, Periodic Table, Unit 3 (Chapter 6) Notes, & Chapter 6 in textbook

Prerequisites:

Answer the following questions using the resources above.

- 1. What happens to the number of valence electrons as you go down a group on the periodic table?
- 2. What happens to the number of valence electrons as you go from left to right across a period?
- 3. What happens to the number of energy levels as you go down a group on the periodic table?
- 4. What happens to the number of energy levels as you go from left to right across a period?
- 5. Draw a Lewis structure for a Cl atom.

- Draw a Bohr diagram for the Cl atom. (Draw above)
- In this lab we will be talking about atomic radius.
 - a. Look up the atomic radius of Cl on Table S:
 - b. Why is it possible to show the atomic radius using a Bohr diagram but not by using a Lewis structure?
 - On your Bohr diagram draw an arrow to represent the radius of the CI atom.
- Define periodic:
- 9. Define trend:

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Table of Contents Geometry Connections Volume 2 Answer Key

- 1. Understanding the eBook Geometry Connections Volume 2 Answer Key
 - The Rise of Digital Reading Geometry Connections Volume 2 Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Geometry Connections Volume 2 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Geometry Connections Volume 2 Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Geometry Connections Volume 2 Answer Key
 - Personalized Recommendations
 - Geometry Connections Volume 2 Answer Key User Reviews and Ratings
 - Geometry Connections Volume 2 Answer Key and Bestseller Lists

- 5. Accessing Geometry Connections Volume 2 Answer Key Free and Paid eBooks
 - Geometry Connections Volume 2 Answer Key Public Domain eBooks
 - Geometry Connections Volume 2 Answer Key eBook Subscription Services
 - Geometry Connections Volume 2 Answer Key Budget-Friendly Options
- 6. Navigating Geometry Connections Volume 2 Answer Key eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Geometry Connections Volume 2 Answer Key Compatibility with Devices
 - Geometry Connections Volume 2 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Geometry Connections Volume 2 Answer Key
 - Highlighting and Note-Taking Geometry Connections Volume 2 Answer Key
 - Interactive Elements Geometry Connections Volume 2 Answer Key
- 8. Staying Engaged with Geometry Connections Volume 2 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Geometry Connections Volume 2 Answer Key
- 9. Balancing eBooks and Physical Books Geometry Connections Volume 2 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Geometry Connections Volume 2 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Geometry Connections Volume 2 Answer Key
 - Setting Reading Goals Geometry Connections Volume 2 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Geometry Connections Volume 2 Answer Key
 - Fact-Checking eBook Content of Geometry Connections Volume 2 Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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