

Name: _____

Date: _____

Empirical vs. Molecular Formulas

Compound	Empirical Formula	Molecular Formula
Water	H_2O	H_2O
Hydrogen Peroxide	HO	H_2O_2
Acetylene	CH	C_2H_2
Benzene	CH	C_6H_6
Glucose	CH_2O	$\text{C}_6\text{H}_{12}\text{O}_6$
Acetic Acid	CH_2O	$\text{C}_2\text{H}_4\text{O}_2$

Answer the following question:

1. What is the difference between an empirical formula and a molecular formula?
2. How can you determine the molecular formula from the empirical formula?

Empirical Formulas And Molecular Formual Answers

Neil D. Jespersen, Alison Hyslop



Empirical Formulas And Molecular Formual Answers:

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Fundamentals of Chemistry: A Modern Introduction Frank Brescia, 2012-12-02 Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry, including the properties of gases, atomic and molecular weights, and the first and second laws of thermodynamics. Chemical equations and chemical arithmetic are also discussed, along with the structure of atoms, chemical periodicity, types of chemical bonds, and condensed states of matter. This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists. Separation and purification are covered in the first chapter, while the following chapters focus on atomic and molecular weights, stoichiometry, the structure of atoms, and types of chemical bonds. The molecular orbital (MO) theory of bonding, galvanic cells, and chemical thermodynamics are considered next. Separate chapters are devoted to MO theory of covalent and metallic bonding, orbital hybridization, intermolecular forces, acids and bases, ionic equilibrium, calculations, and polymers and biochemicals. This monograph is intended for chemistry students.

Chemical Principles Peter Atkins, Loretta Jones, 2007-08 Written for calculus-inclusive general chemistry courses, Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications. Unlike other texts, it begins with a detailed picture of the atom, then builds toward chemistry's frontier, continually demonstrating how to solve problems, think about nature and matter, and visualize chemical concepts as working chemists do. Flexibility in level is crucial and is largely established through clearly labeling, separating in boxes the calculus coverage in the text. Instructors have the option of whether to incorporate calculus in the coverage of topics. The multimedia integration of Chemical Principles is more deeply established than any other text for this course. Through the unique eBook, the comprehensive Chemistry Portal, Living Graph icons that connect the text to the Web, and a complete set of animations, students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding.

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