

# Experiment 7 Report Sheet

## Empirical Formulas

Date \_\_\_\_\_ Lab Sec \_\_\_\_\_ Name \_\_\_\_\_

Indicate whether the following data are for Part B, Part C, or Part D. \_\_\_\_\_

Desk No. \_\_\_\_\_

Identify the original sample \_\_\_\_\_

	Trial 1	Trial 2
1. Mass of crucible and lid (g)	25.636	40.347
2. Mass of crucible, lid, and sample (g)	25.786	40.542
3. Mass of sample (g) 0.150	0.150	0.195
4. Instructor's approval of flame and apparatus	—	—
5. Mass of crucible, lid, and product		
1st mass measurement (g)	25.858	40.654
2nd mass measurement (g)	25.856	40.652
3rd mass measurement (g)	25.859	40.658
6. Final mass of crucible, lid, and product (g)	25.860	40.655
7. Mass of product (g)	0.224g	0.113g
8. Part B. Combination Reaction of Magnesium and Oxygen		
a. Mass of O (g)	0.014g	
b. Mass ratio of Mg to O	2:1	
c. Moles of Mg (mol)	0.00617 mol	
d. Moles of O (mol)	0.00463 mol	
e. Mole ratio of Mg to O	1:1	
f. Consensus empirical formula of magnesium oxide	1 MgO	
g. Percent by mass (%): 67.0 % Mg; 33.0 % O		

### 9. Part C. Decomposition Reaction of a Pure Compound

a. Mass of product = mass CaO (g)	0.224	
b. Mass of CO <sub>2</sub> (g)	0.1263	

e) note

# Empirical Formula Of A Compound Lab Report

**Gary S. Thorpe**



## **Empirical Formula Of A Compound Lab Report:**

**Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

**Basic Laboratory Principles in General Chemistry** Fitzgerald B. Bramwell, 1990      Laboratory Manual Jo A. Beran, 1990      *Bibliography of Scientific and Industrial Reports*, 1947      **Merck's Report** Theodore Weicker, 1910      **Introduction to Chemistry, Laboratory Manual** T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students      Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998      Experimental Chemistry Wilbert Hutton, 1977

Scientific and Technical Aerospace Reports, 1988      Nationalizing Science Alan J. Rocke, 2000-11-08 After looking at the early careers of Wurtz's two mentors Liebig and Jean Baptiste Dumas Rocke describes Wurtz's life and career in the politically complex period leading up to 1853 He then discusses the turning point in Wurtz's intellectual life his conversion to the reformed chemistry of Laurent Gerhardt and Williamson and his efforts to persuade his colleagues of the advantages of

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Bibliography of Scientific and Industrial Reports, 1967      **Government Reports Announcements & Index**, 1988

**Modern Experimental Chemistry** George W. Jr. Latimer, 2012-12-02 Modern Experimental Chemistry provides techniques of qualitative analysis that reinforce experiments on ionic equilibria. This book includes the determination of water in hydrated salts, identification of an organic compound after determining its molecular weight and nonaqueous titration of a salt of a weak acid. The calculation of chemical stoichiometry, calculation of thermodynamic properties by determining the change in equilibrium with temperature and chromium chemistry are also covered. This compilation contains enough experiments for classes which have six hours of laboratory two 3 hour meetings per week to last two semesters. This publication is intended for chemistry students as an introductory manual to chemistry laboratory.

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