

Empirical/Molecular Formula

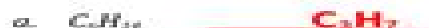
Name: _____

Date: _____

1. What is an empirical formula? How is it different from a molecular formula?

The empirical formula is the lowest whole number ratio of atoms in a compound. The molecular formula shows the true number of atoms in the molecule and is a whole number multiple of the empirical formula.

2. State the empirical formula for each of the following.



3. A sample of a compound is 50.0% sulfur and 50.0% oxygen by mass. What is the empirical formula?

$$S: \frac{50.0 \text{ g S}}{32.07 \text{ g S}} \times \frac{1 \text{ mol S}}{1} = \frac{1.559 \text{ mol S}}{1.559 \text{ mol}} = 1 \text{ S}$$

EF: SO_2

$$O: \frac{50.0 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{3.125 \text{ mol O}}{1.559 \text{ mol}} = 2 \text{ O}$$

4. A sample of a compound is 52.9% aluminum. The rest is oxygen. What is the empirical formula?

$$Al: \frac{52.9 \text{ g Al}}{26.98 \text{ g Al}} \times \frac{1 \text{ mol Al}}{1} = \frac{1.961 \text{ mol Al}}{1.961 \text{ mol Al}} = 1 \text{ Al} \times 2 = 2$$

EF: Al_2O_3

$$O: \frac{47.1 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{2.944 \text{ mol O}}{1.961 \text{ mol}} = 1.5 \times 2 = 3 \text{ O}$$

Ratio must be a whole

5. A compound is 48.65% carbon, 8.11 % hydrogen and 43.24 % oxygen. Determine the empirical formula.

4.05 mol carbon, 8.11 mol H, 2.70 mol O



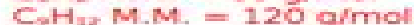
1. Divide by the smallest # to get = $C_{1.5}H_3O_1$

2. Multiply to get a whole number ratio (x2) = $C_3H_6O_2$

6. Vinegar is a dilute solution of acetic acid. The molar mass of acetic acid is 60.0 g/mol and it has an empirical formula of CH_2O . What is the molecular formula of acetic acid?



7. A compound has an empirical formula of C_3H_4 . Which of the following are possible molar masses of the compound? 20 g/mol, 55 g/mol, 80 g/mol, 120 g/mol.



Empirical And Molecular Formula Show Work

Mr. Shaik.Munwar, Dr. Ravi
Ketipally,Mr. Mohammad Mustaq

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style questions Pagination remains the same as GCSE Chemistry so that the two can be used alongside each other

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Table of Contents Empirical And Molecular Formula Show Work

1. Understanding the eBook Empirical And Molecular Formula Show Work
 - The Rise of Digital Reading Empirical And Molecular Formula Show Work
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formula Show Work
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical And Molecular Formula Show Work
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formula Show Work
 - Personalized Recommendations
 - Empirical And Molecular Formula Show Work User Reviews and Ratings
 - Empirical And Molecular Formula Show Work and Bestseller Lists

5. Accessing Empirical And Molecular Formula Show Work Free and Paid eBooks
 - Empirical And Molecular Formula Show Work Public Domain eBooks
 - Empirical And Molecular Formula Show Work eBook Subscription Services
 - Empirical And Molecular Formula Show Work Budget-Friendly Options
6. Navigating Empirical And Molecular Formula Show Work eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical And Molecular Formula Show Work Compatibility with Devices
 - Empirical And Molecular Formula Show Work Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical And Molecular Formula Show Work
 - Highlighting and Note-Taking Empirical And Molecular Formula Show Work
 - Interactive Elements Empirical And Molecular Formula Show Work
8. Staying Engaged with Empirical And Molecular Formula Show Work
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical And Molecular Formula Show Work
9. Balancing eBooks and Physical Books Empirical And Molecular Formula Show Work
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical And Molecular Formula Show Work
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Empirical And Molecular Formula Show Work
 - Setting Reading Goals Empirical And Molecular Formula Show Work
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Empirical And Molecular Formula Show Work
 - Fact-Checking eBook Content of Empirical And Molecular Formula Show Work
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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