

 **K_{sp} and ICE Problems Explained**

Key: K_{sp} = solubility product
 subscripted values

- a. Write out the solubility product value after writing the dissolving solid's equation in the equation:



Initial	K_{sp}	Change	Equilibrium
PbCl_2	1.7×10^{-2}	$-x$	x
Pb^{2+}		$+x$	x
Cl^-		$+2x$	$2x$

- b. Write out the solubility product value if you have a solid for equilibrium with a cell value in the equation in the equation:



Initial	K_{sp}	Change	Equilibrium
PbCl_2	1.7×10^{-2}	$-x$	x
Pb^{2+}		$+x$	x
Cl^-		$+2x$	$2x$

- c. Use what you learned in question b and substitute in concentration values and solve. Remember to make sure that it balances for stoichiometry:



If the problem asks you to find the concentration of a solid, it's usually the only thing that's not given, and it's usually the only thing that's not given.

Initial	K_{sp}	Change	Equilibrium	Initial	Equilibrium
PbCl_2	1.7×10^{-2}	$-x$	x	Pb^{2+}	x
Pb^{2+}		$+x$	x	Cl^-	$2x$
Cl^-		$+2x$	$2x$	PbCl_2	x

Ice Practice Problems Answer Key

Minjie Lin

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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