

- A. 30
- A. 36

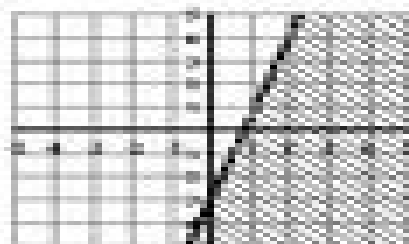
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STAAR CHECKPOINT

- 1 Write an inequality that represents the graph below.



- A $7x + 2y > 6$
- B $7x - 2y > 6$
- C $7x + 2y \leq 6$
- D $7x - 2y \geq 6$

A.28

- 2 Which inequality best represents the table below?

x	y
-5	1
0	3
2	-4
5	0

- A $y > \frac{2}{3}x + 3$
- B $y < \frac{2}{3}x + 3$
- C $y \geq \frac{2}{3}x + 3$
- D $y \leq \frac{2}{3}x + 3$

A.28

- 3 Marcus earns \$10.25 per hour at his part-time job. His weekly pay, y , varies directly with the number of hours, x , he works each week. Which equation could be used to represent this situation?

- A $y = 10.25$
- B $x = 10.25$
- C $x = 10.25y$
- D $y = 10.25x$

A.28

- 4 Jordan is buying gas for his car. The cost of gas varies directly with the number of gallons of gas. If Jordan pays \$8.75 for 5 gallons of gas, how much would it cost for 32 gallons of gas?

- A \$56
- B \$160
- C \$280
- D \$18.29

A.28

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Michael Rajnik



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g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is
reacted with 8.6 g of H_2

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problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical
reaction

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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4 + \text{NaNO}_3$
 which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate
 what is the

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{CO}_2 + \text{I}_2$ a 80.0 grams of
 iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced
 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

7 3 limiting reactant and percent yield problems - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to
 give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3