

Answer Key

PG-1: What is the coefficient of the hydroxide ion for the half reaction when the disproportionation reaction is given the smallest coefficient reduction? $\text{BrO}^-(\text{aq}) \rightarrow \text{Br}^-(\text{aq}) + \text{BrO}_3^-(\text{aq})$ unbalanced



PG-2: What is the balanced half reaction of the species in reaction in basic solution according to the smallest coefficient reduction?



PG-3: What is the coefficient for water when the reaction is balanced in acidic solution? $\text{CrO}_4^{2-}(\text{aq}) + \text{Mn}^{2+}(\text{aq}) \rightarrow \text{Cr}^{3+}(\text{aq}) + \text{MnO}_4^-(\text{aq})$ unbalanced



PG-4: Which reaction occurs in the cell in the presence of standard state conditions?



The standard reduction potential of $\text{Fe}^{3+}/\text{Fe}^{2+}$ is +0.77 V.
The standard reduction potential of Fe^{2+}/Fe is 0 V.



The cell is a concentration cell. The cell potential is 0 V. The cell potential is 0 V. The cell potential is 0 V.

Reduction occurs in the reduction of Fe^{3+} .
For a reduction, the standard potential is the standard potential of the half-reaction. $\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$.
The standard potential is 0.77 V. The standard potential is 0.77 V.
The standard potential is 0.77 V. The standard potential is 0.77 V.

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