

## IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1.  $\text{K} + \text{F}$



2.  $\text{Mg} + \text{I}$



3.  $\text{Be} + \text{S}$



4.  $\text{Na} + \text{O}$



5.  $\text{Al} + \text{Br}$



## COVALENT BONDING

Name \_\_\_\_\_

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:



Note that hydrogen is content with 2, not 8, electrons.

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1.  $\text{H} + \text{H} (\text{H}_2)$



2.  $\text{F} + \text{F} (\text{F}_2)$



3.  $\text{O} + \text{O} (\text{O}_2)$



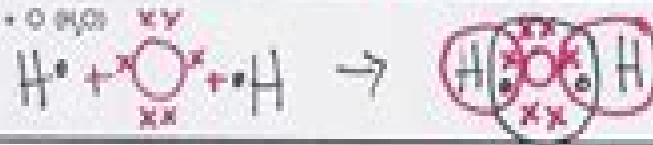
4.  $\text{N} + \text{N} (\text{N}_2)$



5.  $\text{C} + \text{O} (\text{CO})$



6.  $\text{H} + \text{O} (\text{H}_2\text{O})$



# Guided Practice Chapter 8 Covalent Bonding Key

**Robert R. Redfield**



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## **Table of Contents Guided Practice Chapter 8 Covalent Bonding Key**

1. Understanding the eBook Guided Practice Chapter 8 Covalent Bonding Key
  - The Rise of Digital Reading Guided Practice Chapter 8 Covalent Bonding Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Chapter 8 Covalent Bonding Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Practice Chapter 8 Covalent Bonding Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Chapter 8 Covalent Bonding Key
  - Personalized Recommendations
  - Guided Practice Chapter 8 Covalent Bonding Key User Reviews and Ratings
  - Guided Practice Chapter 8 Covalent Bonding Key and Bestseller Lists

- 
5. Accessing Guided Practice Chapter 8 Covalent Bonding Key Free and Paid eBooks
    - Guided Practice Chapter 8 Covalent Bonding Key Public Domain eBooks
    - Guided Practice Chapter 8 Covalent Bonding Key eBook Subscription Services
    - Guided Practice Chapter 8 Covalent Bonding Key Budget-Friendly Options
  6. Navigating Guided Practice Chapter 8 Covalent Bonding Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Guided Practice Chapter 8 Covalent Bonding Key Compatibility with Devices
    - Guided Practice Chapter 8 Covalent Bonding Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Guided Practice Chapter 8 Covalent Bonding Key
    - Highlighting and Note-Taking Guided Practice Chapter 8 Covalent Bonding Key
    - Interactive Elements Guided Practice Chapter 8 Covalent Bonding Key
  8. Staying Engaged with Guided Practice Chapter 8 Covalent Bonding Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Guided Practice Chapter 8 Covalent Bonding Key
  9. Balancing eBooks and Physical Books Guided Practice Chapter 8 Covalent Bonding Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Guided Practice Chapter 8 Covalent Bonding Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Guided Practice Chapter 8 Covalent Bonding Key
    - Setting Reading Goals Guided Practice Chapter 8 Covalent Bonding Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Guided Practice Chapter 8 Covalent Bonding Key
    - Fact-Checking eBook Content of Guided Practice Chapter 8 Covalent Bonding Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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