

Name Key Period _____

Worksheet: Calculating Empirical & Molecular Formulas

1. The empirical formula for the compound having the formula $H_4C_2O_4$ is:

- [A] C_2H_2 [B] CO_2H [C] COH [D] $C_2O_4H_4$ [E] COH_2

2. Calculate the empirical formula of a compound that is 85.6% C and 14.4% H (by mass).

- [A] CH_2 [B] CH [C] C_2H_4 [D] C_2H_6 [E] C_3H_8

$$C: \frac{85.6}{12.01} = 7.12 \quad H: \frac{14.4}{1.01} = 14.25 = \frac{14.25}{7.12} = 2$$

3. A compound is analyzed and found to contain 12.1% carbon, 56.2% oxygen, and 71.7% chlorine (by mass). Calculate the empirical formula of this compound.

- [A] $COCl_2$ [B] CO_2Cl_2 [C] CO_2Cl [D] $COCH$ [E] $COCl$

$$C: \frac{12.1}{12.01} = 1 \quad O: \frac{56.2}{16.00} = 3.51 \quad Cl: \frac{71.7}{35.45} = 2$$

4. A compound contains 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen (by mass). Calculate the empirical formula.

- [A] C_2H_2O [B] CH_2O [C] CH_4O [D] $C_2H_4O_2$ [E] C_2H_4O

$$C: \frac{40.0}{12.01} = 3.33 \quad H: \frac{6.7}{1.01} = 6.63 \quad O: \frac{53.3}{16.00} = 3.33 = \frac{3.33}{3.33} = 1$$

5. A compound contains 25.94% N and 74.06% O (by mass). What is the empirical formula?

$$\frac{25.94g N}{14.01} = \frac{1.85}{1.84} = 1 \times 2 = 2$$

$$\frac{74.06g O}{16.00} = \frac{4.63}{1.84} = 2.5 \times 2 = 5$$



6. Determine the empirical formula of a compound containing 54.2% F and 45.8% S (by mass).

$$\frac{54.2g F}{19.00g/mol} = \frac{2.85}{1.43} mol F = 1.99 \sim 2$$

$$\frac{45.8g S}{32.07g/mol} = \frac{1.43}{1.43} mol S = 1$$



7. A compound has 40.68% carbon, 5.12% hydrogen, and 54.20% oxygen (by mass). Calculate its empirical formula.

$$\frac{40.68g C}{12.01g/mol} = \frac{3.387}{3.387} \quad \frac{5.12g H}{1.01g/mol} = \frac{5.069}{3.387} \quad \frac{54.20g O}{16.00g/mol} = \frac{3.387}{3.387}$$

$$= 1 \times 2 = 2 \quad = 1.496 \times 2 = 3 \quad = 1 \times 2 = 2$$



8. A 7.33-g sample of lanthanum, La, combines with oxygen to give 10.29 g of the oxide. Calculate the empirical formula of this oxide.

$$10.29g - 7.33g La = 2.96g O$$

$$\frac{7.33g La}{138.91g/mol} = .0527 \quad \frac{2.96g O}{16.00g/mol} = .1850 = \frac{.1850}{.0527} = 3.5 \times 2 = 7$$



9. Calculate the molecular formula of a compound with the empirical formula CH_2O and a molar mass of 150 g/mol.

- [A] $C_7H_6O_5$ [B] $C_5H_{10}O_5$ [C] $C_3H_6O_3$ [D] $C_4H_8O_4$ [E] $C_6H_{12}O_6$

$$30 \quad 150 \div 30 = 5$$

Empirical Formula Study Guide With Answer Sheet

BM King



Empirical Formula Study Guide With Answer Sheet:

The book delves into Empirical Formula Study Guide With Answer Sheet. Empirical Formula Study Guide With Answer Sheet is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Empirical Formula Study Guide With Answer Sheet, encompassing both the fundamentals and more intricate discussions.

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