

The Gravimetric Determination of Sulfate

**CHEMISTRY 120 LAB (CHE 120 L) GENERAL CHEMISTRY I
LABORATORY**

**SOUTHERN CONNECTICUT STATE UNIVERSITY DEPARTMENT OF
CHEMISTRY**

Purpose: The purpose of this two part experiment is to quantitatively determine sulfate in an unknown. Using a technique called digestion is used in the gravimetric analysis to determine the amount of sulfate ion, SO_4^{2-} in an unknown sample. Statistical methods will then be applied to the experimental results of the group and of the whole class respectively to determine the success of the experiment.

Introduction: There are two major parts to this lab that both have important significance to the results of the experiment. The first part of this lab is gravimetric analysis. This technique is used to quantitatively determine the amount of chemical in the unknown solution (201). Results for this genre of measurement are generally expressed in terms of percentage of a solution. Gravimetric determination is determined by precipitating the unknown as an insoluble and then weighing the dried precipitate. To attain a good sample one must dry the precipitate to remove

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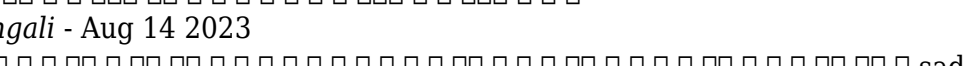
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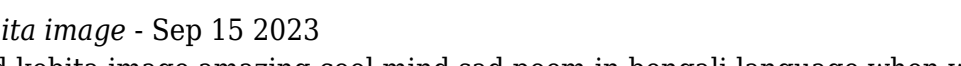
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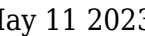
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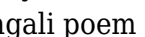
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