

# ISC Class 12 Chemistry Question Paper Solution 2017

## CHEMISTRY THEORY (PAPER-1)

### Part I (20 marks)

Answer all questions

#### Question 1

- (a) Fill in the blanks by choosing the appropriate word/words from those given in the brackets: [5]  
(iodoform, acetaldehyde, positive, greater, acidic, acetone, disaccharide, negative, increases, glucose, decreases, chloroform, polysaccharide, lactose, lesser, basic, cationic hydrolysis, anionic hydrolysis)
- (i) Calcium acetate on heating gives \_\_\_\_\_ which gives \_\_\_\_\_ on heating with iodine and sodium hydroxide solution.
- (ii) On dilution of a solution, its specific conductance \_\_\_\_\_ while its equivalent conductance \_\_\_\_\_.
- (iii) Sucrose is a \_\_\_\_\_ and yields upon hydrolysis, a mixture of \_\_\_\_\_ and fructose.
- (iv) More \_\_\_\_\_ is the standard reduction potential of a substance, the \_\_\_\_\_ is its ability to displace hydrogen from acids.
- (v) An aqueous solution of  $\text{CH}_3\text{COONa}$  is \_\_\_\_\_ due to \_\_\_\_\_.
- (b) Complete the following statements by selecting the correct alternative from the choices given: [5]
- (i) In a face centered cubic lattice, atom (A) occupies the corner positions and atom (B) occupies the face centre positions. If one atom of (B) is missing from one of the face centered points, the formula of the compound is:
- (1)  $\text{A}_2\text{B}_5$   
(2)  $\text{A}_2\text{B}_3$   
(3)  $\text{AB}_2$   
(4)  $\text{A}_2\text{B}$
- (ii) The half-life period of a first order reaction is 20 minutes. The time required for the concentration of the reactant to change from 0.16 M to 0.02M is:
- (1) 80 minutes  
(2) 60 minutes  
(3) 40 minutes  
(4) 20 minutes
- (iii) For a spontaneous reaction  $\Delta G^\circ$  and  $E^\circ$  cell will be respectively:

# Isc Chemistry Paper

**Padhraic Smyth**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Isc Chemistry Paper**

1. Understanding the eBook Isc Chemistry Paper
  - The Rise of Digital Reading Isc Chemistry Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Isc Chemistry Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Isc Chemistry Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Isc Chemistry Paper
  - Personalized Recommendations
  - Isc Chemistry Paper User Reviews and Ratings

- Isc Chemistry Paper and Bestseller Lists
- 5. Accessing Isc Chemistry Paper Free and Paid eBooks
  - Isc Chemistry Paper Public Domain eBooks
  - Isc Chemistry Paper eBook Subscription Services
  - Isc Chemistry Paper Budget-Friendly Options
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  - Isc Chemistry Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Isc Chemistry Paper
  - Highlighting and Note-Taking Isc Chemistry Paper
  - Interactive Elements Isc Chemistry Paper
- 8. Staying Engaged with Isc Chemistry Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Isc Chemistry Paper
- 9. Balancing eBooks and Physical Books Isc Chemistry Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Isc Chemistry Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Isc Chemistry Paper
  - Setting Reading Goals Isc Chemistry Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Isc Chemistry Paper
  - Fact-Checking eBook Content of Isc Chemistry Paper
  - Distinguishing Credible Sources

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  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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