| Name | |
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Periodic Table Trend Activities:

Purpose: The Periodic Table is a useful tool that helps us predict properties of various elements. In this activity, we will look at 3 properties:

- 1. Atomic radius: The distance from the atomic nucleus to the outmost occupied orbital in a atom
- Electronegativity: The ability of an atom to attract a bonded pair of electrons from another atom
- 3. Ionization Energy: The energy needed to remove an electron from a neutral gaseous atom
- 4. Electron Affinity: The energy released when an electron in added to a neutral gaseous atom

Procedure: Click on the <u>Cool Periodic Table link</u> (click on the element symbol). Record the atomic radius, electronegativity, and first ionization energy for each element in the table below:

http://www.lvnchburg.net/hhs/chemistry/trends/

Durtue:

| Element | Atomic Radius | Electronegativity | 1 st Ionization Energy |
|---------|---------------|-------------------|-----------------------------------|
| Li | | | |
| Be | | | |
| В | | |), |
| Na | | | |
| Mg | | | T. |
| · Al | | | |
| C | | | |
| Si | | | |
| N | | | |
| P | | | T |

Discussion:

- What appears to be the trend in atomic radius as you move from left to right in a row?
- 2. What appears to be the trend in atomic radius as you move down a column?

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Yicheng Fang

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