

# BASIC CHEMISTRY

## CHEMISTRY

- ✓ Medieval Latin "Alchimia"/ Arabic "Alkimya", the chemistry
- ✓ The study of the composition and structure of the changes in matter
- ✓ It is often called the central science because knowledge in this field is a prerequisite for all the other physical and biological sciences.

## HISTORICAL DEVELOPMENT OF THE ATOMIC STRUCTURE

### ATOM

- ✓ It is the smallest possible unit into which matter can be divided, while still maintaining its properties
- ✓ Building blocks of matter
- ✓ It has protons (+), neutrons (0), and electrons (-)

1. **Democritus** was a greek philosopher who first introduced the concept of atoms. He believed that matter is comprises of tiny indivisible particles known as atomos.
2. **Dalton's Atomic Theory**
  - a. All matter is made up of indivisible particles known as atoms.
  - b. All atoms of the same elements are alike.
  - c. Compounds are formed when atoms of different elements combine in certain proportions.
  - d. atoms are only rearranged but neither created nor destroyed in chemical reactions.
3. **Law of Definite Proportion** – a given compound always contains the same elements in exactly the same proportion by mass.

$H_2O$  - ratio of two hydrogen atom with one oxygen atom only. If it becomes ratio of two hydrogen atom with two oxygen atom, it is not already water.

4. **Law of Multiple Proportions** – when two elements can combine to form more than one type of compound, the masses of one element that combine with a fixed mass of the other elements are in a ratio of small whole numbers.



5. **Law of Conservation of Mass** states that matter can neither be created nor destroyed. Matter is just transformed.
6. The discovery of the different sub-atomic particles:
  - i. **J.J. Thompson** determined the charge to mass ratio of the electron using a cathode ray tube.

# Example Basic Chemistry Study Guide

**M Lipman**



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**Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

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