

ANSWERS

Section III: Problems (Answers to odds in back of the text book)

Practice Problems:

Chem Practice Problems:

1. Ammonia is a strong household cleaner. What is the atomic mass of one molecule?
17.04amu
2. The double replacement ...
3. A bottle of lead II sulfate contains 158.1g of the compound. How many moles of the compound are present?
0.5213 mol PbSO₄
4. Find the
5. Determine the number of formula units that are in 1.48 mol NaF
 8.91×10^{23} units NaOH
6. How many formula ...
7. How many atoms are in 1.25 mol of oxygen gas?
 1.27×10^{24} atoms O
8. If you burned 6.10×10^{24} ...
9. A container with a volume of 101L contains how many moles of argon gas at STP?
4.51 mol Ar
10. Find the percent composition ...
11. A sample of an unknown compound with a mass of 0.847g has the following composition 50.51% F & 49.49% Fe (m/m). When the compound is decomposed into its elements, what is the contributing mass of each? Write the decomposition equation after reassembling the empirical formula.
.428g F & .419g Fe

 $\frac{.0225\text{mol F}}{.00751\text{mol}} = 3$ & $\frac{.00751\text{mol Fe}}{.00751\text{mol}} = 1$; $2\text{FeF}_3 \rightarrow 2\text{Fe} + 3\text{F}_2$
12. Find the molecular formula ...
13. A 50.51g sample of made of phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39g of phosphorus atoms were produced.
 - a. What is the empirical formula of the compound?
 $11.39/30.97 = \frac{.368\text{mol P}}{.368\text{mol}}$ & $(50.51\text{g} - 11.39\text{g})/35.45\text{g} = \frac{.551\text{mol Cl}}{.368\text{mol}}$

= 1 P & 1.5 Cl (double to make whole #s) P₂Cl₃

General Chemistry Homework Guide

O García



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Table of Contents General Chemistry Homework Guide

1. Understanding the eBook General Chemistry Homework Guide
 - The Rise of Digital Reading General Chemistry Homework Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Homework Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Homework Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Homework Guide
 - Personalized Recommendations
 - General Chemistry Homework Guide User Reviews and Ratings
 - General Chemistry Homework Guide and Bestseller Lists
5. Accessing General Chemistry Homework Guide Free and Paid eBooks
 - General Chemistry Homework Guide Public Domain eBooks
 - General Chemistry Homework Guide eBook Subscription Services
 - General Chemistry Homework Guide Budget-Friendly Options

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 - General Chemistry Homework Guide Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of General Chemistry Homework Guide
 - Highlighting and Note-Taking General Chemistry Homework Guide
 - Interactive Elements General Chemistry Homework Guide
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Homework Guide
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10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Homework Guide
 - Setting Reading Goals General Chemistry Homework Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Homework Guide
 - Fact-Checking eBook Content of General Chemistry Homework Guide
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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