

Explain how you identified the limiting and excess reactant based only on your qualitative observations (not quantitative measurements or calculations). Type your answer into the box. Use complete sentences.

According to observation the initial color of bromocresol purple is yellow in color after adding the NaOH, it turns violet. After Changing the color yellow from purple indicates acid is completely consumed so excess of base addition changes the color into violet.

Table C. Calculations to Determine Theoretical Yields

<i>Description</i>	<i>Calculated Value</i>
Volume of NaOH solution	22.20 mL
Initial moles of NaOH	0.002451 moles
Initial moles of H_3PO_4	0.002602 moles
Mass of H_2O produced if NaOH runs out first	0.04414 g
Mass of H_2O produced if H_3PO_4 runs out first	1.265 g
Theoretical yield of H_2O (in grams)	1.265 g
Theoretical yield of Na_3PO_4 (in grams)	0.4267 g

Table D. Identification of Limiting and Excess Reactants from Calculations

<i>Description</i>	<i>Reactant Name or Formula</i>
Limiting Reactant	Phosphoric Acid
Excess Reactant	Sodium Hydroxide

Explain how you identified the limiting and excess reactant based on your calculations. Be specific. Type your answer into the box. Use complete sentences.

According to our calculation the reacting equivalent of H_3PO_4 is less than NaOH. Therefore H_3PO_4 is limiting reactant so product formation depends on H_3PO_4 because phosphoric acid is completely consumed.

Formal Lab Report On Limiting Reactant

A Loxley



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