Empirical Formula, Molecular Formula and Hydrates Worksheet

1. Write the Empirical Formula for Each of the Following:

a. P₄O₆

e. C₆H₈O₆ _____

b. C₆H₉

f. C₁₀H₂₂

e. CH2OHCH2OH_____

g. Cu₂C₂O₄ _____

d. BrCl-

h. Hg₂F₂

Write the empirical formula for each of the following (show your work):

A compound composed of: 72% iron (Fe) and 27.6% oxygen (O) by mass.

- A compound composed of: 9.93% carbon (C), 58.6% chlorine (C1), and 31.4% fluorine (F).
 (This compound is commonly known as Freen)
- c. A compound composed of: 0.556g carbon (C) and 0.0933g hydrogen (H).
- Write the molecular formula for each of the following: (amu atomic mass units, for our purposes it is equivalent to molecular mass of the compound.
 - A compound with a molecular mass of 70.0 amu and an empirical formula of CH₂.
 - A compound with a molecular mass of 46.0 amu and an empirical formula of NO₂.
- 4. Can the molecular formula of a compound ever be the same as the empirical formula for the compound? Explain you answer.

Empirical Molecular Formula Worksheet Answers

Annelies Wilder-Smith

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Empirical Molecular Formula Worksheet Answers Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the ability of words has be much more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Empirical Molecular Formula Worksheet Answers**, a literary masterpiece that delves deep in to the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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