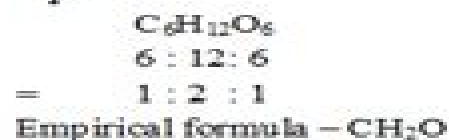
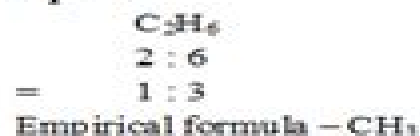


Empirical formula: It shows the simplest whole number ratio between the atoms of the elements in the compound.

Example 1:



Example 2:



Example 3:



Practice:

1. C_4H_8
2. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
3. N_2O_4

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Calculate the empirical formula of the compounds having the following percentage composition:

Example 1:

C-40%
H-6.7%
O-53.3%

Sol:

| Element | % | % ÷ At. Wt. | Simplest ratio | simplest whole number ratio |
|-------------------|------|--------------------------|-------------------------|-----------------------------|
| Carbon | 40 | $40 \div 12 = 3.33$ | $\frac{3.33}{3.33} = 1$ | 1 C |
| Hydrogen | 6.7 | $6.7 \div 1$ | $\frac{6.7}{3.33} = 2$ | 2 H |
| Oxygen | 53.3 | $53.3 \div 16$ | $\frac{3.33}{3.33} = 1$ | 1 O |
| Empirical Formula | | = CH₂O | | |

Empirical Formula Practice

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



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