

Gas Laws Worksheet

Boyle, Charles, Pressure and Combined Gas Laws

Boyle's Law Problems:

$$P_1V_1 = P_2V_2$$

$$1 \text{ atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 790 mm Hg at constant temperature. What is the new volume?
2. A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L. What is the pressure in the container if the temperature remains constant?
3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?
4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

Charles' Law Problems:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$1 \text{ atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

5. A container containing 5.00 L of a gas is collected at 100 K and then allowed to expand to 20.0 L. What must the new temperature be in order to maintain the same pressure (as required by Charles' Law)?
6. A gas occupies 900.0 mL at a temperature of 27.0 °C. What is the volume at 132.0 °C?
7. If 15.0 liters of neon at 25.0 °C is allowed to expand to 45.0 liters, what must the new temperature be to maintain constant pressure?

Pressure (Lussac) Law Problems:

$$P_1/T_1 = P_2/T_2$$

8. Determine the pressure change when a constant volume of gas is heated from 20.0°C to 30.0°C.
9. If a gas in a closed container is pressurized from 15 atm to 16 atm and its original temperature was 25°C, determine its final temperature.
10. The temperature of a sample of gas in a sealed container at 30.0 kPa is increased from -100°C to 1000°C. Determine the final temperature.

Mixed Review Gases Answers

J Ma



Mixed Review Gases Answers:

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Table of Contents Mixed Review Gases Answers

1. Understanding the eBook Mixed Review Gases Answers
 - The Rise of Digital Reading Mixed Review Gases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Review Gases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Review Gases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Review Gases Answers
 - Personalized Recommendations
 - Mixed Review Gases Answers User Reviews and Ratings
 - Mixed Review Gases Answers and Bestseller Lists

5. Accessing Mixed Review Gases Answers Free and Paid eBooks
 - Mixed Review Gases Answers Public Domain eBooks
 - Mixed Review Gases Answers eBook Subscription Services
 - Mixed Review Gases Answers Budget-Friendly Options
6. Navigating Mixed Review Gases Answers eBook Formats
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 - Mixed Review Gases Answers Compatibility with Devices
 - Mixed Review Gases Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Review Gases Answers
 - Highlighting and Note-Taking Mixed Review Gases Answers
 - Interactive Elements Mixed Review Gases Answers
8. Staying Engaged with Mixed Review Gases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Review Gases Answers
9. Balancing eBooks and Physical Books Mixed Review Gases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Review Gases Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mixed Review Gases Answers
 - Setting Reading Goals Mixed Review Gases Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mixed Review Gases Answers
 - Fact-Checking eBook Content of Mixed Review Gases Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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