Thermochemistry and Gas Laws Unit Review

Energy Math

 How much energy is released when a brick with a mass of 329g is cooled from 80°C to 50°C? The specific heat capacity of brick is 0.84 1/g°C.

What is the mass of a hotdog that releases 3600 J of energy as it is cooled from 95°C to 45 °C?
The specific heat capacity of a hotdog is 1.5 J/g°C.

 What is the specific heat capacity of a 350 g object that rises in temperature 15°C when 55,900 J of energy are added?

 45,920 J of energy are added to a 365g block of copper originally at 35°C. What will the new temperature of the copper be? The specific heat capacity of copper is 0.385 ½g°C.

$$\Delta t = \frac{45920J}{(365)(0.585)(5)} = \frac{326°C}{135°C} = 77 - 35°C + 35°C$$

 89,290 J of energy are removed from a 670 g block of aluminum originally at 125°C. What will the new temperature of the aluminum be? The specific heat capacity of aluminum is 0.902 U/g°C.

$$\Delta + \frac{-892905}{(6705)(0.900360)} = \frac{-14800}{+1250} = \frac{70 - 1250}{+1250}$$

Energy And Thermochemistry Review Guide Key

James E. Brady, Fred Senese

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