

The Mole Study Guide

Mass Composition: What's your element's % of mass after the reaction? Molar mass to obtain your value!

10. There are 3.75 moles of CO_2 .

11. There are _____

12. There are _____

13. There are _____

14. There are 2.5 moles of _____

15. There are _____

16. There are _____

17. There are _____

18. There are 4.5 moles of _____

19. There are _____

20. There are _____

21. There are 3.5 moles of _____

22. There are _____

23. There are _____

24. There are 2.5 moles of _____

25. There are _____

26. There are _____

Name _____ Period _____ Date _____

The Mole Study Guide

Fill in the Blank boxes for the image below.



Write the Empirical Formula for each compound. Use subscripts to show the relative ratio of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_8\text{H}_{16}\text{O}_8$	
$\text{C}_{12}\text{H}_{24}\text{O}_{12}$	
$\text{C}_{16}\text{H}_{32}\text{O}_{16}$	

Calculate the Molar Mass for the following compounds.

10. H_2O

11. H_2SO_4

12. H_2O_2

13. H_2SO_3

Calculate the Percent Composition for each element in the following compound.

10. H_2O

11. H_2SO_4

12. H_2O_2

13. H_2SO_3

14. H_2O_2

15. H_2SO_3

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 40.00% Hg, 21.87% S, 3.75% N, and 34.38% O. The molar mass for Selenium is 78.96 g/mol.

a. If you have 100 g of Selenium, how many grams of Hg, N, and O do you have?

b. Determine the Empirical Formula for Selenium.

c. Determine the Molecular Formula for Selenium.

Mole Study Guide

Judd E. Hollander



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