

Stoichiometry and Mole Ratios

Mole ratio

For the equation $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the coefficients tell us the following:

1. For every _____ moles of H_2O_2 reacted _____ moles of H_2O are produced.
 $2 \text{ mol H}_2\text{O}_2 = 2 \text{ mol H}_2\text{O}$
2. For every _____ moles of H_2O_2 reacted _____ mole of O_2 is produced
 $2 \text{ mol H}_2\text{O}_2 = 1 \text{ mol O}_2$
3. For every _____ moles of H_2O produced _____ mole of O_2 is produced.
 $2 \text{ mol H}_2\text{O} = 1 \text{ mol O}_2$

Examples:

1. Given the following balanced chemical equation:



- a. How many moles of O_2 are needed to completely react with 2.25 moles of C_3H_8 ?
- b. How many moles of C_3H_8 are required to produce 1.5 moles of H_2O ?

2. How many moles of KOH are required to make 3.0 moles of K_2SO_4 ?



Mole Ratios Worksheet Model 1 A Chemical Reaction

Michael Rajnik



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