

THE MOLE AND AVOGADRO'S NUMBER

Name KEY

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules.

How many molecules are in the quantities below?

1. 2.0 moles	$\frac{2.0 \cancel{\text{mole}}}{1 \cancel{\text{mole}}}$	$\frac{6.02 \times 10^{23} \text{ molecules}}{1 \cancel{\text{mole}}} = \boxed{1.20 \times 10^{24} \text{ molecules}}$
2. 1.5 moles	$\frac{1.5 \cancel{\text{mole}}}{1 \cancel{\text{mole}}}$	$\frac{6.02 \times 10^{23} \text{ molecules}}{1 \cancel{\text{mole}}} = \boxed{9.03 \times 10^{23} \text{ molecules}}$
3. 0.75 mole	$\frac{0.75 \cancel{\text{mole}}}{1 \cancel{\text{mole}}}$	$\frac{6.02 \times 10^{23} \text{ molecules}}{1 \cancel{\text{mole}}} = \boxed{4.52 \times 10^{23} \text{ molecules}}$
4. 15 moles	$\frac{15 \cancel{\text{mole}}}{1 \cancel{\text{mole}}}$	$\frac{6.02 \times 10^{23} \text{ molecules}}{1 \cancel{\text{mole}}} = \boxed{9.03 \times 10^{24} \text{ molecules}}$
5. 0.35 mole	$\frac{0.35 \cancel{\text{mole}}}{1 \cancel{\text{mole}}}$	$\frac{6.02 \times 10^{23} \text{ molecules}}{1 \cancel{\text{mole}}} = \boxed{2.11 \times 10^{23} \text{ molecules}}$

How many moles are in the number of molecules below?

1. 6.02×10^{23}	$\boxed{1 \text{ mole}}$ (By definition!)
2. 1.204×10^{24}	$\frac{1.204 \times 10^{24} \text{ molecules}}{6.02 \times 10^{23} \text{ molecules}} = \boxed{2.00 \text{ moles}}$
3. 1.5×10^{20}	$\frac{1.5 \times 10^{20} \cancel{\text{mole}}}{6.02 \times 10^{23} \cancel{\text{molecules}}} = \boxed{2.49 \times 10^{-4} \text{ mole}}$
4. 3.4×10^{26}	$\frac{3.4 \times 10^{26} \cancel{\text{mole}}}{6.02 \times 10^{23} \cancel{\text{molecules}}} = \boxed{565 \text{ moles}}$
5. 7.5×10^{19}	$\frac{7.5 \times 10^{19} \cancel{\text{mole}}}{6.02 \times 10^{23} \cancel{\text{molecules}}} = \boxed{1.24 \times 10^{-4} \text{ moles}}$

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SJ Ball



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