

# Mole Conversions Practice Answer Key

Molar Mass – Calculate the Molar Mass for the following.

- What is the molar mass of Magnesium metal?

$$24.305 \text{ g/mol}$$

- Calculate the molar mass of NaCl

$$\begin{aligned} \text{Na} &: 22.990 \times 1 = 22.990 \\ \text{Cl} &: 35.453 \times 1 = 35.453 \end{aligned} \Rightarrow 58.443 \text{ g/mol}$$

- What is the molar mass of NaOH?

$$\begin{aligned} \text{Na} &: 1 \times 22.990 = 22.990 \\ \text{O} &: 1 \times 15.999 = 15.999 \\ \text{H} &: 1 \times 1.008 = 1.008 \end{aligned} \Rightarrow 39.997 \text{ g/mol}$$

Conversions – Calculate the following the conversions. Use your Mole Map to help you!

- Determine the number of atoms in 0.50 moles He.

$$0.5 \text{ mol He} \times \frac{6.02 \times 10^{23} \text{ atoms He}}{1 \text{ mol He}} = 3.01 \times 10^{23} \text{ atoms He}$$

- Determine the number of moles in 80.00 g O<sub>2</sub>

$$0.2 \times 15.999 = 31.998 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol} = 31.998 \text{ g}$$

- Determine the volume of 0.24 mol of NO<sub>2</sub>.

$$0.24 \text{ mol NO}_2 \times \frac{22.4 \text{ L NO}_2}{1 \text{ mol NO}_2} = 5.376 \text{ L NO}_2$$

$$80.00 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 25.00 \text{ g O}_2$$

- Determine the number of moles in  $7.5 \times 10^{19}$  formula units of KNO<sub>3</sub>

$$7.5 \times 10^{19} \text{ f.u. KNO}_3 \times \frac{1 \text{ mol KNO}_3}{6.02 \times 10^{23} \text{ f.u. KNO}_3} = 1.2458 \times 10^{-4} \text{ f.u. KNO}_3$$

- Determine the number of moles in 202.22g KNO<sub>3</sub>

$$\begin{aligned} \text{K} &: 1 \times 39.098 = 39.098 \\ \text{N} &: 1 \times 14.007 = 14.007 \\ \text{O} &: 3 \times 15.999 = 47.997 \end{aligned} \Rightarrow 101.102 \text{ g/mol} \Rightarrow 1 \text{ mol} = 101.102 \text{ g}$$

- Determine the number of moles in 32.5L of N<sub>2</sub> at STP

$$32.5 \text{ L N}_2 \times \frac{1 \text{ mol N}_2}{22.4 \text{ L N}_2} = 1.451 \text{ mol N}_2$$

$$202.22 \text{ g KNO}_3 \times \frac{1 \text{ mol KNO}_3}{101.102 \text{ g KNO}_3} = 2.00 \text{ mol KNO}_3$$

- Determine the number of moles in  $3.4 \times 10^{26}$  formula units of NaCl.

$$3.4 \times 10^{26} \text{ f.u. NaCl} \times \frac{1 \text{ mol NaCl}}{6.02 \times 10^{23} \text{ f.u. NaCl}} = 569.784 \text{ f.u. NaCl}$$

- Determine the moles in 36.5L of carbon monoxide

$$36.5 \text{ L CO} \times \frac{1 \text{ mol CO}}{22.4 \text{ L CO}} = 1.629 \text{ mol CO}$$

# Mole 8 1 Answer Key

**Disha Experts**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

## **Mole 8 1 Answer Key:**

## **Mole 8 1 Answer Key Book Review: Unveiling the Magic of Language**

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Mole 8 1 Answer Key**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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### **Table of Contents Mole 8 1 Answer Key**

1. Understanding the eBook Mole 8 1 Answer Key
  - The Rise of Digital Reading Mole 8 1 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole 8 1 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole 8 1 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole 8 1 Answer Key
  - Personalized Recommendations
  - Mole 8 1 Answer Key User Reviews and Ratings
  - Mole 8 1 Answer Key and Bestseller Lists

5. Accessing Mole 8 1 Answer Key Free and Paid eBooks
  - Mole 8 1 Answer Key Public Domain eBooks
  - Mole 8 1 Answer Key eBook Subscription Services
  - Mole 8 1 Answer Key Budget-Friendly Options
6. Navigating Mole 8 1 Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Mole 8 1 Answer Key Compatibility with Devices
  - Mole 8 1 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole 8 1 Answer Key
  - Highlighting and Note-Taking Mole 8 1 Answer Key
  - Interactive Elements Mole 8 1 Answer Key
8. Staying Engaged with Mole 8 1 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole 8 1 Answer Key
9. Balancing eBooks and Physical Books Mole 8 1 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole 8 1 Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mole 8 1 Answer Key
  - Setting Reading Goals Mole 8 1 Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mole 8 1 Answer Key
  - Fact-Checking eBook Content of Mole 8 1 Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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