Name_______Date:

The Mole Webquest

Use the following website to answer this set of questions:

http://antoine.frostburg.edu/chem/senese/101/moles/faq/why-use-moles.shtml

1) A mole of anything is how many? (give the number):

 6.02×10^{22}

 Why is it that different amounts of things can still equal one mole? (think about the weight of a dozen elephants vs a dozen eggs)

The number of grams in a mole is different from substance to substance.

3) Why do we want to use the concept of moles?

You often want to know how many molecules you have in a sample of a substance.

With moles, you count the number of molecules in the sample by weighing it.

Once we know the number of moles we can convert to the number of:
 moles or grams and vice versa.

5) How many grams of water are in one mole of water?

1 mole water = 18 g water

6) How many molecules of water are in one mole of water?

 6.02×10^{23}

Use the following website to answer the next set of questions: https://chemfiesta.wordpress.com/2015/02/06/mole-calculations-3/

Define molar mass.

The mass of a given substance (chemical element or chemical compound is in g) divided by its amount of substance in (mol).

8) What is the mass of one mole of oxygen? 16.0 grams per mole

8) Find the molar mass of NH3NO3. (Show your work)

Ammonium nitrate has two nitrogen atoms (2 x 14.0 g/mol), four hydrogen atoms (4 x 1.0 g/mol), and three oxygen atoms (3 x 16.0 g/mol). This makes the molar mass 80.0 g/mol.

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R Bogdan



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