

Dilutions – Practice Problems

Instructions: Use the equation for dilutions to complete the problems below.

- 1) A student obtains 0.30 liters of 2.0 M silver nitrate solution. She dilutes the solution until it has a volume of 2.6 liters. What is the **molarity** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(2.0) (0.30) = M_2 (2.6)$$

$$(0.6) = M_2 (2.6)$$

$$M_2 = 0.23 \text{ M}$$

- 2) Before an experiment, a scientist obtains 2.45 L of 1.25 M sulfuric acid (H_2SO_4) and dilutes it until the molarity drops to 0.80 M. What is the **volume** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(1.25) (2.45) = (0.8) V_2$$

$$(3.06) = (0.8) V_2$$

$$V_2 = 3.83 \text{ L}$$

- 3) A solution of hydrochloric acid is diluted, bringing the volume from 2.0 L to 8.5 L. If the solution has a final molarity of 1.4 M, what was the **molarity before** dilution?

$$M_1 V_1 = M_2 V_2$$

$$M_1 (2.0) = (1.4) (8.5)$$

$$(250.0) = (2.0) V_2$$

$$V_2 = 125.0 \text{ mL}$$

- 4) If a student adds 250 mL of water to a 300 mL solution of NaCl that has a concentration of 3.0 M, what will be the **molarity** of the final solution?

$$M_1 V_1 = M_2 V_2$$

$$(3.0) (300) = M_2 (550)$$

$$(900) = M_2 (550)$$

$$M_2 = 1.64 \text{ M}$$



- 5) A scientist adds 1.2 liters of water to a 5.0 liter solution, causing the molarity to become 0.4 M. What was the **molarity** of the **initial solution** before the dilution?

$$M_1 V_1 = M_2 V_2$$

$$M_1 (5.0) = (0.4) (7.2)$$

$$M_1 (5.0) = 2.88$$

$$M_1 = 0.576 \text{ M}$$

Molarity And Dilution Practice Problems Key

**Morris Hein, Scott Pattison, Susan
Arena, Leo R. Best**



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Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed

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