

# Chemical Equations and Reactions

Quiz  
CH 8  
Section #1

## Section Quiz: Describing Chemical Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- B 1. Which of the following is true regarding the law of conservation of mass?
- It dictates that the number of molecules on each side of a chemical equation must be the same.
  - It dictates that the number of atoms of each element must be the same on both sides of a chemical equation.
  - It states that the mass of the reactants must remain constant in order for a chemical reaction to proceed.
  - It does not apply to chemical reactions.
- D 2. Which of the following may indicate that a chemical reaction has occurred?
- release of energy as light
  - a color change
  - gas bubble formation
  - All of the above
- D 3. Which of the following is a product in the reaction described by the word equation below?
- $$\text{iron} + \text{copper(II) sulfate} \rightarrow \text{iron(II) sulfate} + \text{copper}$$
- iron(II) sulfate
  - copper
  - iron
  - Both (a) and (b)
- C 4. What does  $\text{CH}_4$  represent in the balanced chemical equation shown below?
- $$\text{C} + 2\text{H}_2 \rightarrow \text{CH}_4$$
- one atom of  $\text{CH}_4$
  - five atoms of  $\text{CH}_4$
  - one molecule of  $\text{CH}_4$
  - 1 g of  $\text{CH}_4$
- C 5. Which of the following is not true regarding a properly written chemical equation?
- It must be balanced.
  - It must contain proper chemical formulas.
  - It must show the formation of a precipitate or a gas.
  - It should represent known data.

# Modern Chemistry Study Guide Answers Every Chapter

**RJ Alexander**



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## **Table of Contents Modern Chemistry Study Guide Answers Every Chapter**

1. Understanding the eBook Modern Chemistry Study Guide Answers Every Chapter
  - The Rise of Digital Reading Modern Chemistry Study Guide Answers Every Chapter
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Study Guide Answers Every Chapter
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Study Guide Answers Every Chapter
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Study Guide Answers Every Chapter
  - Personalized Recommendations
  - Modern Chemistry Study Guide Answers Every Chapter User Reviews and Ratings
  - Modern Chemistry Study Guide Answers Every Chapter and Bestseller Lists
5. Accessing Modern Chemistry Study Guide Answers Every Chapter Free and Paid eBooks
  - Modern Chemistry Study Guide Answers Every Chapter Public Domain eBooks
  - Modern Chemistry Study Guide Answers Every Chapter eBook Subscription Services
  - Modern Chemistry Study Guide Answers Every Chapter Budget-Friendly Options
6. Navigating Modern Chemistry Study Guide Answers Every Chapter eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Study Guide Answers Every Chapter Compatibility with Devices
  - Modern Chemistry Study Guide Answers Every Chapter Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Study Guide Answers Every Chapter
  - Highlighting and Note-Taking Modern Chemistry Study Guide Answers Every Chapter
  - Interactive Elements Modern Chemistry Study Guide Answers Every Chapter
8. Staying Engaged with Modern Chemistry Study Guide Answers Every Chapter

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Modern Chemistry Study Guide Answers Every Chapter
- 9. Balancing eBooks and Physical Books Modern Chemistry Study Guide Answers Every Chapter
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Study Guide Answers Every Chapter
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Study Guide Answers Every Chapter
  - Setting Reading Goals Modern Chemistry Study Guide Answers Every Chapter
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Study Guide Answers Every Chapter
  - Fact-Checking eBook Content of Modern Chemistry Study Guide Answers Every Chapter
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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