

# Modern Chemistry<sup>®</sup>

## Study Guide



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 time the total pressure  $p_a = x_a p_{\text{total}}$  mole fraction  $x_a = \frac{\text{moles of substance a}}{\text{total moles}}$   
 a is often denoted by  $x_a$  and

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 by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction  
 mole fraction is not dependent on the temperature

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 mass  $C_6H_6 = 78 \text{ g/mol}$  hence  $C_6H_6$ 's number of moles  $\frac{30}{78} = 0.3846 \text{ mol}$  molar mass of carbon tet  $CCl_4 = 154 \text{ g/mol}$

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 g of water calculate the molarity and molality of the solution given  $n = 14$   $h = 1$   $c = 12$   $o = 16$  given mass of solute urea 11.11 g mass  
 of solvent water 100 g 0.1 kg to find molarity of solution solution

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 of  $CH_3OH = \frac{0.1722}{0.1722 + 0.073} = 0.703$  to solve more examples on mole fraction formula and practice more  
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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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