

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H}-3$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

# Modern Chemistry Seven Section 1 Study Guide

**Ying Liu**



## **Modern Chemistry Seven Section 1 Study Guide:**

## Adopting the Beat of Term: An Mental Symphony within **Modern Chemistry Seven Section 1 Study Guide**

In a world consumed by monitors and the ceaseless chatter of immediate interaction, the melodic elegance and psychological symphony developed by the prepared word often fade in to the background, eclipsed by the relentless noise and disturbances that permeate our lives. But, set within the pages of **Modern Chemistry Seven Section 1 Study Guide** a marvelous fictional value filled with natural feelings, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, that fascinating masterpiece conducts visitors on an emotional trip, well unraveling the concealed tunes and profound influence resonating within each cautiously constructed phrase. Within the depths of this emotional evaluation, we will investigate the book is central harmonies, analyze their enthralling publishing type, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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### **Table of Contents Modern Chemistry Seven Section 1 Study Guide**

1. Understanding the eBook Modern Chemistry Seven Section 1 Study Guide
  - The Rise of Digital Reading Modern Chemistry Seven Section 1 Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Seven Section 1 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Seven Section 1 Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Seven Section 1 Study Guide
  - Personalized Recommendations

- Modern Chemistry Seven Section 1 Study Guide User Reviews and Ratings
- Modern Chemistry Seven Section 1 Study Guide and Bestseller Lists
- 5. Accessing Modern Chemistry Seven Section 1 Study Guide Free and Paid eBooks
  - Modern Chemistry Seven Section 1 Study Guide Public Domain eBooks
  - Modern Chemistry Seven Section 1 Study Guide eBook Subscription Services
  - Modern Chemistry Seven Section 1 Study Guide Budget-Friendly Options
- 6. Navigating Modern Chemistry Seven Section 1 Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Seven Section 1 Study Guide Compatibility with Devices
  - Modern Chemistry Seven Section 1 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Seven Section 1 Study Guide
  - Highlighting and Note-Taking Modern Chemistry Seven Section 1 Study Guide
  - Interactive Elements Modern Chemistry Seven Section 1 Study Guide
- 8. Staying Engaged with Modern Chemistry Seven Section 1 Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Seven Section 1 Study Guide
- 9. Balancing eBooks and Physical Books Modern Chemistry Seven Section 1 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Seven Section 1 Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Seven Section 1 Study Guide
  - Setting Reading Goals Modern Chemistry Seven Section 1 Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Seven Section 1 Study Guide
  - Fact-Checking eBook Content of Modern Chemistry Seven Section 1 Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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