



Name		Date	Class	
Modern Chemistry	CHAPTER 7			

#### **HOMEWORK 7-1**

(pp. 263-266)

#### VOCABULARY

Write true or faire for each statement.

- A chemical formula indicates the relative number of molecules of each kind in a chemical compound.
- 2. Monoatomic ions are ions formed from a single atom.
- 3. Monoatomic anions of an element are identified by the element's name.
- 4. A binary compound is made of two or more different elements.
- 5 The Stock system is a system for naming chemical elements and their ions.

#### SKILL BUILDER

Explain what each chemical formula tells you about the composition of the compound. The first one is done for you.

#### NuiNO.

The subscript 3 refers to the number of oxygen atoms. There are one atom of sodium and one atom of nitrogen for every three atoms of oxygen.

AgCL

BlackOffly,

(NH<sub>e</sub>)<sub>2</sub>SO<sub>2</sub>

#### STANDARDIZED TEST PREP

Circle the letter of the best answer.

- Which statement is true?
  - a. All main-group elements are able to form ions.
  - b. The total numbers of positive charges and negative charges must be equal in a binary compound.
  - c. The nonmetals of groups 15, 16, and 17 lose electrons to form cations.
  - d. Elements in the d block form ions of only one charge.
- 2. Why do many main-group elements lose or gain electrons to form ions?
  - The elements are unstable and therefore lose electrons.
  - b. Their tendency is to maximize potential energy.
  - c. The elements form covalent bonds, sharing their electrons with other atoms.
  - d. Their tendency is to form a complete outermost octet in a noble-gas configuration.

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**Wolfgang Guggemos** 

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