#### Chapter 8: Particle Physics

#### Outline the results of the gold leaf experiments

- Alpha particles fired at gold metal foil
- In a vacuum.
- Most pass straight through/undeflected;
- Some deflected:
- Very few reflected/deflected through large angles;
- Atom mainly empty space;
- Most did not get near enough to any matter to be deflected;
- Only a few particles came close enough to charge to be affected;
- Nucleus has mass much greater than alpha particle mass to cause deflection;
- 10. Nucleus contains most of the mass;
- 11. Nucleus very small/tiny;
- 12. Nucleus charged/positive;
- Alpha particles are deflected by a charged nucleus that has a very small diameter compared to that of the atom rather than the charge being distributed throughout the atom;
- 14. most of the mass of the atom is concentrated in the nucleus rather than distributed throughout the atom;

## Explain how the observations from the alpha particle scattering experiment led to a new model of the atom

- 1. Most alpha particles pass through undeflected OR some deflected through a small angle;
- 2. A very small number are deflected through an angle greater than 90 degrees;
- This suggests that the alpha particles are deflected by a charged nucleus that has a very small diameter compared to that of the atom rather than the charge being distributed throughout the atom;
- and that most of the mass of the atom is concentrated in the nucleus rather than distributed throughout the atom;

#### Role of E and M fields in particle detector

- Electric field accelerates particles
- Direction of deflection/force indicates sign of charge;
- 3. a=EQ/m;
- Magnetic fields provide a centripetal force";
- Direction of force/curvature indicates sign of charge;
- Momentum from radius;
- r= p/8Q;

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**Ariel Salleh** 

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