

17: Molecular Geometry - VSEPR

Key Structure Terms

- Valence Shell:** Electrons in the outermost shell that are involved in bonding.
- Lewis Structure:** A 2D representation of a molecule and its bonds.
- Lone Pair:** Pair of electrons not being shared in a bond.
- Bonding Pair:** Pair of electrons used to make a bond. Both atoms sharing the electrons can "count" them in their valence shell.
- Single bond:** One pair of shared electrons.
- Double bond:** Two pairs of shared electrons. Shorter and stronger than a single bond.
- Triple bond:** Three pairs of shared electrons. Shorter and stronger than a double bond.
- Valence Shell Electron Pair Repulsion Theory (VSEPR):** Bonds and lone pairs are electrons, all electrons are negative, negative things repel other negative things, therefore all bonds and lone pairs arrange themselves in 3D as far away from each other as possible.
- Electron Geometry:** 3D structure of a molecule determined by counting the electron regions around a central atom (bonds and lone pairs).
- Electron Region:** Each bond (single, double or triple) and lone pair count as "1" electron region.
- Molecular Geometry:** 3D structure determined by the atoms bonded to the central atom.
- Ligand:** Atoms bonded to the central atom.

VSEPR Theory and Geometry

Electron geometry is determined by looking at the number of electron regions around the central atom.

Molecular geometry is determined by looking at the number of atoms bonded to the central atom (ligand) and the number of lone pairs around the central atom.

Electronic Geometry Mnemonic: Linear(2) – Trigonal Planar(3) – Tetrahedral(4) – Trigonal Bipyramidal(5) – Octahedral(6) = "Long DIP To TRIBE Overseas."

Examples:
















Electron geometry: Tetrahedron
Molecular geometry: Tetrahedron



Electron geometry: Tetrahedron
Molecular geometry: Trigonal pyramidal

Molecular Geometries

A = central atom; X = ligands; E = lone pairs

Electron Regions	Molecular Formula	Name	Shape
2	AX_2 (BeCl_2 , CO_2)	Linear	
3	AX_3 (BF_3 , CO_3^{2-} , NO_3^- , SO_3)	Trigonal Planar	
	AX_2E (NO_2^+ , SO_2 , O_3)	Bent	
4	AX_4 (CH_4 , NH_4^+ , PO_4^{3-} , SO_4^{2-} , ClO_4^-)	Tetrahedron	
	AX_3E (NH_3 , H_3O^+ , PCl_3 , SO_3^{2-})	Trigonal pyramidal	
	AX_2E_2 (H_2O , ClO_2 , OF_2 , SCl_2)	Bent	
5	AX_5 (PCl_5)	Trigonal bipyramidal	
	AX_4E (SF_4 , SCl_5)	See-saw	
	AX_3E_2 (ClF_3 , ICl_3)	T-shaped	
	AX_2E_3 (XeF_2 , I_3^-)	Linear	
6	AX_6 (SF_6 , PCl_6^-)	Octahedron	
	AX_5E (BrF_5 , IF_5)	Square pyramidal	
	AX_4E_2 (XeF_4)	Square planar	

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Molecular Structure And Bonding Cheat Sheet

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Table of Contents Molecular Structure And Bonding Cheat Sheet

1. Understanding the eBook Molecular Structure And Bonding Cheat Sheet
 - The Rise of Digital Reading Molecular Structure And Bonding Cheat Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Structure And Bonding Cheat Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molecular Structure And Bonding Cheat Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Structure And Bonding Cheat Sheet
 - Personalized Recommendations
 - Molecular Structure And Bonding Cheat Sheet User Reviews and Ratings
 - Molecular Structure And Bonding Cheat Sheet and Bestseller Lists
5. Accessing Molecular Structure And Bonding Cheat Sheet Free and Paid eBooks
 - Molecular Structure And Bonding Cheat Sheet Public Domain eBooks
 - Molecular Structure And Bonding Cheat Sheet eBook Subscription Services
 - Molecular Structure And Bonding Cheat Sheet Budget-Friendly Options

6. Navigating Molecular Structure And Bonding Cheat Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Molecular Structure And Bonding Cheat Sheet Compatibility with Devices
 - Molecular Structure And Bonding Cheat Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molecular Structure And Bonding Cheat Sheet
 - Highlighting and Note-Taking Molecular Structure And Bonding Cheat Sheet
 - Interactive Elements Molecular Structure And Bonding Cheat Sheet
8. Staying Engaged with Molecular Structure And Bonding Cheat Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molecular Structure And Bonding Cheat Sheet
9. Balancing eBooks and Physical Books Molecular Structure And Bonding Cheat Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molecular Structure And Bonding Cheat Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molecular Structure And Bonding Cheat Sheet
 - Setting Reading Goals Molecular Structure And Bonding Cheat Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molecular Structure And Bonding Cheat Sheet
 - Fact-Checking eBook Content of Molecular Structure And Bonding Cheat Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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