Key

# Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

#### Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

#### Model 1 - Lewis Structures

Lows Structures	H <sub>2</sub> CO 3 electron domains (3 bonding, 0 nonbonding)	3-D Molecular Shape
2. Ref. #06#	BeF <sub>2</sub> 2 electron domains (2 bonding, 0 nonbonding)	2 Tetrakedra
3.04	CH, 4 electron domains (4 bonding, 0 nonbonding)	o go Tetrakedra
5. H <sub>2</sub> O	NH <sub>3</sub> 4 electron domains (3 bonding, 1 nonbonding)	800 pyrami
·.o,	H <sub>3</sub> O 4 electron domains (2 bonding, 2 nonbonding)	a Bent
Lone pair - ·-	CO <sub>2</sub> 2 electron domains (2 bonding, 0 nonbonding)	O=O=O

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Mary Millhollon, Richard H. Langley

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