

Name \_\_\_\_\_

## Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	No of electron domains/lig	Electron domain geometry	VSEPR Molecular geometry	Hybridization of central atom	Polar ?
$\text{CO}_2$	$\text{O}=\text{C}=\text{O}$ 24 valence 24 2 charge clouds	0	Linear 180°	Linear 	$\text{sp}$	NP
$\text{BF}_3$	$\text{F}-\text{B}-\text{F}$ 24 valence 32 3 charge clouds	0	Trigonal planar 120°		$\text{sp}^2$	NP
$\text{SnCl}_2$	$\text{Cl}-\text{Sn}-\text{Cl}$ 16 valence 24 2 charge clouds	3	Bent <120°		$\text{sp}^2$	P
$\text{BF}_4^-$	$\left[ \text{F}-\text{B}-\text{F} \right]^-$ 24 valence 40 4 charge clouds	0	Tetrahedral 109.5°		$\text{sp}^3$	P
$\text{AsH}_3$	$\text{H}-\text{As}-\text{H}$ 8 valence 26 3 charge clouds	4	Trigonal pyramidal 109.5°		$\text{sp}^3$	P
$\text{H}_2\text{S}$	$\text{H}-\text{S}-\text{H}$ 8 valence 26 2 charge clouds	2	Bent 109.5°		$\text{sp}^3$	NP
$\text{NH}_2^-$	$\left[ \text{H}-\text{N}-\text{H} \right]^-$ 8 valence 26 2 charge clouds	2	Bent 109.5°		$\text{sp}^3$	P
$\text{POF}_3$	$\text{F}-\text{P}-\text{Cl}$ 10 valence 32 4 charge clouds	0	Trigonal bipyramidal 120° + 90°		$\text{sp}^3\text{d}$	P

# Molecular Geometry And Polarity Lab Answers

**M Woodhall**



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