

CHAPTER 10 The Mole

Section 1 Measuring Matter

page 320–324

Practice Problems

pages 323–324

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol Zn} \times \frac{6.02 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H₂O).

$$11.5 \text{ mol H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules H}_2\text{O}$$

3. Silver nitrate (AgNO₃) is used to make several different silver halides used in photographic films. How many formula units of AgNO₃ are there in 3.25 mol AgNO₃?

$$3.25 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units AgNO}_3$$

4. **Challenge** Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule, O₂.

$$5.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} \\ \times \frac{2 \text{ O atoms}}{\text{molecule O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?
 - a. 5.75×10^{24} atoms Al

$$5.75 \times 10^{24} \text{ atoms Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms Al}} \\ = 9.55 \text{ mol Al}$$

- b. 2.50×10^{20} atoms Fe

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$$2.50 \times 10^{20} \text{ atoms Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. **Challenge** Identify the representative particle for each formula and convert the given number of representative particles to moles.

- a. 3.75×10^{24} CO₂

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules CO}_2 \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 6.23 \text{ mol CO}_2$$

- b. 3.58×10^{23} ZnCl₂

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula units ZnCl}_2 \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula units ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

Section 1 Review

page 324

7. **Explain** why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. **State** the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number (6.02×10^{23}) of representative particles.

Mole Solution Manual

Daniel X. B. Chen



Mole Solution Manual:

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 - The Rise of Digital Reading Mole Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Solution Manual
 - Personalized Recommendations
 - Mole Solution Manual User Reviews and Ratings
 - Mole Solution Manual and Bestseller Lists
5. Accessing Mole Solution Manual Free and Paid eBooks
 - Mole Solution Manual Public Domain eBooks
 - Mole Solution Manual eBook Subscription Services
 - Mole Solution Manual Budget-Friendly Options
6. Navigating Mole Solution Manual eBook Formats

- ePub, PDF, MOBI, and More
- Mole Solution Manual Compatibility with Devices
- Mole Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole Solution Manual
 - Highlighting and Note-Taking Mole Solution Manual
 - Interactive Elements Mole Solution Manual
- 8. Staying Engaged with Mole Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Solution Manual
- 9. Balancing eBooks and Physical Books Mole Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Solution Manual
 - Setting Reading Goals Mole Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Solution Manual
 - Fact-Checking eBook Content of Mole Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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