

Module Overview

What are the coefficients in a chemical equation for combustion?

What?

A chemical equation represents what we see when we observe a reaction and predict products from reactants. Coefficients in formulas before them are necessary to conserve mass during a chemical reaction. Typically, when we balance the chemical equation we start by conserving individual products. However, in real life, the reaction represented by an equation occurs as unsegmented, constant volume. Thus, if we have very large molecules (C_{10}H_8) we expect to have all such molecules in the equation. Even so, we balance chemical equations as if they were made of individual molecules, often by beginning on one side with 1 and seeing how well we can balance the different types of chemical reaction we are considering.

Problem 1 – A Chemical Reaction



1. Consider the reaction in Problem 1.

a. What are the coefficients for each of the following substances in the reaction?



b. Does your coefficient balance or otherwise the equation in Problem 1?

2. Consider each reaction below as a chemical reaction in Problem 1.

a. Halving the amount of reaction consumed and produced twice.

b. Doubling the rate of C_2H_2 or O_2 or CO_2 . Balance the reaction the fastest while coefficient possible.

	C_2H_2 Consumed	O_2 Consumed	CO_2 Produced	Ratio $\text{C}_2\text{H}_2/\text{O}_2/\text{CO}_2$ (possible)
How a single product from many reactants? (What coefficient would be consumed or produced?)				
If the reaction consumed less than double reactants, how many molecules would be consumed or produced?				
If the reaction consumed $\frac{1}{100}$ times fewer reactants, would be consumed or produced?				

Mole Ratio Pogil Worksheet Answers

Peipei Pang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Mole Ratio Pogil Worksheet Answers:

Reviewing **Mole Ratio Pogil Worksheet Answers**: Unlocking the Spellbinding Force of Linguistics

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