

Name: KEY

Date: \_\_\_\_\_

Complete the table by determining the number of protons, neutrons, electrons, valence electrons & ionic charge.

neutral  
ion  
isotope

	Protons	Neutrons	Electrons	Valence e-	Ionic charge
Cl	17	18	17	7	0
Al <sup>3+</sup>	13	14	10	3	+3
Gallium- 71	31	39	31	3	0

What element has the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^2$ ? Sn - Tin

Write the orbital notation for each of the following elements:

- Magnesium  $1s \uparrow\downarrow 2s \uparrow\downarrow 2p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow 3s \uparrow\downarrow$
- Germanium  $1s \uparrow\downarrow 2s \uparrow\downarrow 2p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow 3s \uparrow\downarrow 3p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow 4s \uparrow\downarrow 3d \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow 4p \uparrow\downarrow$

Write the electron configurations for the following elements:

- Mn  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^5$
- Nb  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^3$

Write the Nobel gas notation for the following elements:

- Cadmium  $[Kr] 5s^2 4d^{10}$
- Tellurium  $[Kr] 5s^2 4d^{10} 5p^4$

Give the quantum numbers for the LAST electron in each element:

- Strontium  
 $n=5$   
 $l=0$   
 $m_l=0$   
 $s=-1/2$
- Nickel  
 $n=4$   
 $l=2$   
 $m_l=0$   
 $s=-1/2$

# Mole Webquest 2 Answer Key

**JL Elias**



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