

Mole to Grams, Grams to Moles Conversions Worksheet

What are the molecular weights of the following compounds? (all masses must be to nearest hundredth)

- | | |
|----------------------|--|
| 1) NaOH | 2) H ₃ PO ₄ |
| 3) H ₂ O | 4) Mn ₂ Se ₇ |
| 5) MgCl ₂ | 6) (NH ₄) ₂ SO ₄ |

There are three definitions (equalities) of mole. They are:

1 mole = 6.02×10^{23} particles

1 mole = molar mass (could be atomic mass from periodic table or molecular mass)

1 mole = 22.4 L of a gas at STP (You do not need to worry about this yet)

Each definition can be written as a set of two conversion factors. They are:

1 mole = molar mass(g) can be written as $\left(\frac{1 \text{ mole}}{\text{molar mass (g)}} \right)$ OR $\left(\frac{\text{molar mass (g)}}{1 \text{ mole}} \right)$

1 mole = 6.02×10^{23} particles can be written as $\left(\frac{1 \text{ mole}}{6.02 \times 10^{23}} \right)$ OR $\left(\frac{6.02 \times 10^{23}}{1 \text{ mole}} \right)$

Solve the following:

- 1) How many moles are in 15 grams of lithium? (molar mass of lithium is 6.94 g/mole)

$$15 \text{ grams} \times \frac{1 \text{ mole}}{6.94 \text{ grams}} = 2.1614 \text{ moles lithium} = \boxed{2.2 \text{ moles Li}}$$

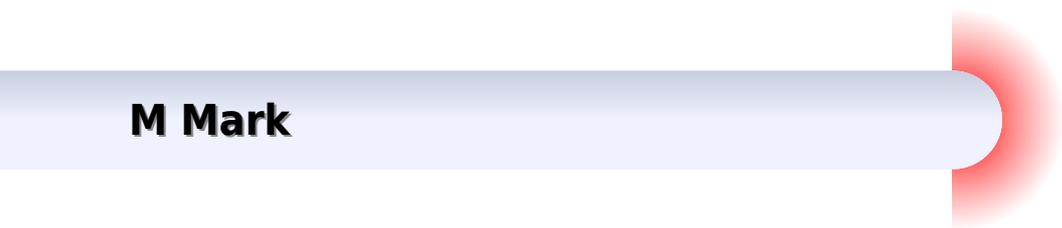
- 2) How many grams are in 2.4 moles of sulfur? (molar mass of sulfur is 32.07 g/ mole)

$$2.4 \text{ moles} \times \frac{32.07 \text{ grams}}{1 \text{ mole}} = 76.97 \text{ grams sulfur} = \boxed{77 \text{ g Sulfur}}$$

- 3) How many moles are in 22 grams of argon?
- 4) How many grams are in 88.1 moles of magnesium?
- 5) How many moles are in 2.3 grams of phosphorus?

Mole Calculation Practice Worksheet Chemfiesta Answers

M Mark



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