



Unit: Stoichiometry **Answer key** "Mole-Mole Calculations" - WS #1

Directions: For each of the following problems, show your balanced chemical equation and all of your calculations.

1. How many moles of potassium oxide (K_2O) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:



$$1.52 \text{ mol } K \times \frac{2 \text{ mol } K_2O}{4 \text{ mol } K} = 0.760 \text{ mol } K_2O$$

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:



$$0.25 \text{ mol } Cu_2S \times \frac{2 \text{ mol } Cu}{1 \text{ mol } Cu_2S} = 0.50 \text{ mol } Cu$$

2. How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:



$$3.75 \text{ mol } Na_2SO_4 \times \frac{2 \text{ mol } Na}{1 \text{ mol } Na_2SO_4} = 7.50 \text{ mol } Na$$

3. How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:



$$1.6 \text{ mol } Fe_2O_3 \times \frac{3 \text{ mol } CaO}{1 \text{ mol } Fe_2O_3} = 4.8 \text{ mol } CaO$$

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: **magnesium chlorate \rightarrow magnesium chloride + oxygen**

Formula equation: **$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$**
(must be balanced!)

$$5.12 \text{ mol } Mg(ClO_3)_2 \times \frac{3 \text{ mol } O_2}{1 \text{ mol } Mg(ClO_3)_2} = 15.4 \text{ mol } O_2$$

5. How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: **chromium + lead (IV) nitrate \rightarrow chromium (III) nitrate + lead**

Formula equation: **$4Cr + 3Pb(NO_3)_4 \rightarrow 4Cr(NO_3)_3 + 3Pb$**
(must be balanced!)

$$0.85 \text{ mol } Pb(NO_3)_4 \times \frac{4 \text{ mol } Cr(NO_3)_3}{3 \text{ mol } Pb(NO_3)_4} = 1.1 \text{ mol } Cr(NO_3)_3$$

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Table of Contents **Mole 2 Answer Key With Work**

1. Understanding the eBook **Mole 2 Answer Key With Work**
 - The Rise of Digital Reading **Mole 2 Answer Key With Work**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Mole 2 Answer Key With Work**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Mole 2 Answer Key With Work**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Mole 2 Answer Key With Work**
 - Personalized Recommendations
 - **Mole 2 Answer Key With Work** User Reviews and Ratings
 - **Mole 2 Answer Key With Work** and Bestseller Lists

5. Accessing Mole 2 Answer Key With Work Free and Paid eBooks
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 - Mole 2 Answer Key With Work eBook Subscription Services
 - Mole 2 Answer Key With Work Budget-Friendly Options
6. Navigating Mole 2 Answer Key With Work eBook Formats
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 - Mole 2 Answer Key With Work Compatibility with Devices
 - Mole 2 Answer Key With Work Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mole 2 Answer Key With Work
 - Highlighting and Note-Taking Mole 2 Answer Key With Work
 - Interactive Elements Mole 2 Answer Key With Work
8. Staying Engaged with Mole 2 Answer Key With Work
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole 2 Answer Key With Work
9. Balancing eBooks and Physical Books Mole 2 Answer Key With Work
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole 2 Answer Key With Work
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mole 2 Answer Key With Work
 - Setting Reading Goals Mole 2 Answer Key With Work
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mole 2 Answer Key With Work
 - Fact-Checking eBook Content of Mole 2 Answer Key With Work
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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