

Moz

Molarity, Molality, & Solutions Practice

Molarity $M = \frac{\text{moles}}{L}$

- Find the M of 1 mole of salt in 2.5 L. of water.

$$\frac{1 \text{ mole}}{2.5 L} = \boxed{.4 M}$$

- Find the M of 2.3 moles of salt in 1.14L. of water.

$$\frac{2.3 \text{ moles}}{1.14} = \boxed{2 M}$$

- Find the M of 1.9 moles of sodium acetate in 12.4L. of water.

$$\frac{1.9}{12.4} = \boxed{.15 M}$$

- How many moles of potassium carbonate are needed to make 500 mL. of a 1.25 M solution?

$$1.25 M = \frac{\text{moles}}{.5 L} \quad (.5)(1.25) = \text{moles} \quad \boxed{0.625 \text{ moles}}$$

- How many liters of 2.5 M solution can be made using 1.8 moles of lithium bromide?

$$2.5 M = \frac{1.8 \text{ moles}}{V} \quad \frac{(2.5)(V)}{2.5} = \frac{1.8}{2.5} \quad \boxed{0.72 L}$$

- What is the concentration of a 450 mL. solution that contains 200 grams of iron (II) chloride?

$$\frac{200 \text{ g FeCl}_2}{112.15 \text{ g}} \left| \frac{1 \text{ mole FeCl}_2}{112.15 \text{ g}} \right| = 1.77 \text{ moles} \quad \frac{1.77}{.45 L} = \boxed{3.93 M}$$

- How many grams of ammonium sulfate are needed to make 0.75L. of solution at a concentration of 4.0 M?

$$4 M = \frac{\text{moles}}{0.75} = 3 \text{ moles} \quad \frac{3 \text{ moles } (NH_4)_2SO_4}{132.17 \text{ g}} \left| \frac{1 \text{ mole } (NH_4)_2SO_4}{132.17 \text{ g}} \right| = \boxed{396.5 \text{ g}}$$

Molality

- Calculate the molality of a solution containing 2.89g of NaCl dissolved in 0.159 L. of water.

$$\frac{2.89 \text{ g NaCl}}{58.44 \text{ g}} \left| \frac{1 \text{ mole}}{58.44 \text{ g}} \right| = 0.049 \text{ mol} \quad \frac{0.049}{0.159} = \boxed{.31 m}$$

- Calculate the molality of a solution containing 1.80 mol KCl in 16.0 mol of H₂O

$$\frac{16 \text{ mol H}_2\text{O}}{1 \text{ mol}} \left| \frac{18 \text{ g}}{1 \text{ mol}} \right| = 288.32 \text{ g} = \text{mol} \quad \frac{1.8 \text{ mol}}{288.32 \text{ g}} = \boxed{6.25 m}$$

Molarity Practice Unit Lesson Key

Hussin A.Rothana



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Table of Contents Molarity Practice Unit Lesson Key

1. Understanding the eBook Molarity Practice Unit Lesson Key
 - The Rise of Digital Reading Molarity Practice Unit Lesson Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Practice Unit Lesson Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Practice Unit Lesson Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Practice Unit Lesson Key

- Personalized Recommendations
- Molarity Practice Unit Lesson Key User Reviews and Ratings
- Molarity Practice Unit Lesson Key and Bestseller Lists
- 5. Accessing Molarity Practice Unit Lesson Key Free and Paid eBooks
 - Molarity Practice Unit Lesson Key Public Domain eBooks
 - Molarity Practice Unit Lesson Key eBook Subscription Services
 - Molarity Practice Unit Lesson Key Budget-Friendly Options
- 6. Navigating Molarity Practice Unit Lesson Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Practice Unit Lesson Key Compatibility with Devices
 - Molarity Practice Unit Lesson Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Practice Unit Lesson Key
 - Highlighting and Note-Taking Molarity Practice Unit Lesson Key
 - Interactive Elements Molarity Practice Unit Lesson Key
- 8. Staying Engaged with Molarity Practice Unit Lesson Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Practice Unit Lesson Key
- 9. Balancing eBooks and Physical Books Molarity Practice Unit Lesson Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Practice Unit Lesson Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity Practice Unit Lesson Key
 - Setting Reading Goals Molarity Practice Unit Lesson Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity Practice Unit Lesson Key

- Fact-Checking eBook Content of Molarity Practice Unit Lesson Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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