

Dilutions – Practice Problems

Instructions: Use the equation for dilutions to complete the problems below.

- 1) A student obtains 0.30 liters of 2.0 M silver nitrate solution. She dilutes the solution until it has a volume of 2.6 liters. What is the **molarity** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(2.0) (0.30) = M_2 (2.6)$$

$$(0.6) = M_2 (2.6)$$

$$M_2 = 0.23 \text{ M}$$

- 2) Before an experiment, a scientist obtains 2.45 L of 1.25 M sulfuric acid (H_2SO_4) and dilutes it until the molarity drops to 0.80 M. What is the **volume** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(1.25) (2.45) = (0.8) V_2$$

$$(3.06) = (0.8) V_2$$

$$V_2 = 3.83 \text{ L}$$

- 3) A solution of hydrochloric acid is diluted, bringing the volume from 2.0 L to 8.5 L. If the solution has a final molarity of 1.4 M, what was the **molarity before** dilution?

$$M_1 V_1 = M_2 V_2$$

$$M_1 (2.0) = (1.4) (8.5)$$

$$(250.0) = (2.0) V_2$$

$$V_2 = 125.0 \text{ mL}$$

- 4) If a student adds 250 mL of water to a 300 mL solution of NaCl that has a concentration of 3.0 M, what will be the **molarity** of the final solution?

$$M_1 V_1 = M_2 V_2$$

$$(3.0) (300) = M_2 (550)$$

$$(900) = M_2 (550)$$

$$M_2 = 1.64 \text{ M}$$



- 5) A scientist adds 1.2 liters of water to a 5.0 liter solution, causing the molarity to become 0.4 M. What was the **molarity** of the **initial solution** before the dilution?

$$M_1 V_1 = M_2 V_2$$

$$M_1 (5.0) = (0.4) (7.2)$$

$$M_1 (5.0) = 2.88$$

$$M_1 = 0.576 \text{ M}$$

Molarity Notes And Worksheet

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Table of Contents Molarity Notes And Worksheet

1. Understanding the eBook Molarity Notes And Worksheet
 - The Rise of Digital Reading Molarity Notes And Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Notes And Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Notes And Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Notes And Worksheet
 - Personalized Recommendations

- Molarity Notes And Worksheet User Reviews and Ratings
- Molarity Notes And Worksheet and Bestseller Lists
- 5. Accessing Molarity Notes And Worksheet Free and Paid eBooks
 - Molarity Notes And Worksheet Public Domain eBooks
 - Molarity Notes And Worksheet eBook Subscription Services
 - Molarity Notes And Worksheet Budget-Friendly Options
- 6. Navigating Molarity Notes And Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Notes And Worksheet Compatibility with Devices
 - Molarity Notes And Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Notes And Worksheet
 - Highlighting and Note-Taking Molarity Notes And Worksheet
 - Interactive Elements Molarity Notes And Worksheet
- 8. Staying Engaged with Molarity Notes And Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Notes And Worksheet
- 9. Balancing eBooks and Physical Books Molarity Notes And Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Notes And Worksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity Notes And Worksheet
 - Setting Reading Goals Molarity Notes And Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity Notes And Worksheet
 - Fact-Checking eBook Content of Molarity Notes And Worksheet

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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