

3. Refer to the data table in Question 2.

- a. How do the reduced ratios in the last column compare to the coefficients in the reaction shown in Model 1?

The reduced ratios are the same as the coefficients in the chemical reaction

- b. Use mathematical concepts to explain how your answer in part a is possible.

If we multiply all the numbers by the same value, the ratio will not change.

4. Even 538 is a small number of molecules to use in a reaction. Typically, chemists use much larger numbers of molecules. (Recall that one mole is equal to 6.02×10^{23} particles.) Consider each situation below as it relates to the reaction in Model 1: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$.

- a. Calculate the amount of reactants consumed and products made.

- b. Record the ratio of N_2 to H_2 to NH_3 . Reduce the ratio to the lowest whole number possible.

	N_2 Consumed	H_2 Consumed	NH_3 Produced	Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$
If the reaction occurred 6.02×10^{23} times, how many molecules would be consumed or produced?	6.02×10^{23}	18.06×10^{23}	12.04×10^{23}	1:3:2
How many <i>moles</i> of each substance would be consumed or produced?	1	3	2	1:3:2

5. Refer to the data table in Question 4.

- a. How do the reduced ratios in the last column compare to the coefficients in the reaction in Model 1?

The reduced ratios are the same as the coefficients in the chemical reaction

- b. Use mathematical concepts to explain how your answer in part a is possible.

Multiplying by the number of particles in a mole, or dividing by three, doesn't change the ratio.

6. The ratio obtained from the coefficients in a balanced chemical equation is called the *mole ratio*.

- a. What is the mole ratio for the reaction in Model 1?

1:3:2

- b. Explain why this ratio is called the mole ratio.

The ratio reflects the number of moles consumed and produced

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