

Example:

25. Three grams of carbon and 15 g of oxygen are placed in a closed vessel and heated to 1000°C. What mass of CO_2 is formed?

$$25 \text{ g } \text{C} \times \frac{1 \text{ mol C}}{12 \text{ g C}} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol C}} \times \frac{44 \text{ g CO}_2}{1 \text{ mol CO}_2} = \boxed{74.4 \text{ g CO}_2}$$

26. When 10 g of aluminum reacts with oxygen, 25 g of Al_2O_3 is formed. How much oxygen was used?



$$10 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{3 \text{ mol O}_2}{2 \text{ mol Al}} \times \frac{32 \text{ g O}_2}{1 \text{ mol O}_2} = 60.6 \text{ g O}_2 \text{ (theoretical)}$$

$$\% \text{ yield} = \frac{25 \text{ g Al}_2\text{O}_3}{60.6 \text{ g Al}_2\text{O}_3} \times 100 = \boxed{73.8\%}$$

27. In the reaction of problem 26, what is the limiting reagent when 10 g of Al and 10 g of O_2 are used?

$$10 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{3 \text{ mol O}_2}{2 \text{ mol Al}} \times \frac{32 \text{ g O}_2}{1 \text{ mol O}_2} = 3.0 \text{ g O}_2$$

$$25 \text{ g Al}_2\text{O}_3 \times \frac{1 \text{ mol Al}_2\text{O}_3}{102 \text{ g Al}_2\text{O}_3} \times \frac{2 \text{ mol Al}}{1 \text{ mol Al}_2\text{O}_3} \times \frac{27 \text{ g Al}}{1 \text{ mol Al}} = \boxed{2.78 \text{ g Al}}$$

Limiting Reagent = Al

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Kentaro Ito

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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