

# Chemical Bonding

## SECTION 3

**SHORT ANSWER** Answer the following questions in the space provided.

1. **A** The notation for sodium chloride, NaCl, stands for one  
(a) formula unit. (c) crystal.  
(b) molecule. (d) atom.
2. **D** In a crystal of an ionic compound, each cation is surrounded by a number of  
(a) molecules. (c) dipoles.  
(b) positive ions. (d) negative ions.
3. **B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is  
(a) higher in potential energy. (c) equal in potential energy.  
(b) lower in potential energy. (d) unstable.
4. **B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?  
(a) Compound A is not an ionic compound.  
(b) It will be more difficult to break the bonds in compound A than those in compound B.  
(c) Compound B has larger crystals than compound A.  
(d) Compound A has larger crystals than compound B.
5. **B** The forces of attraction between molecules in a molecular compound are generally  
(a) stronger than the attractive forces among formula units in ionic bonding.  
(b) weaker than the attractive forces among formula units in ionic bonding.  
(c) approximately equal to the attractive forces among formula units in ionic bonding.  
(d) equal to zero.
6. Describe the force that holds two ions together in an ionic bond.

**Very strong overall attraction between positive and negative charges**

7. What type of energy best represents the strength of an ionic bond?

**Lattice energy**

8. What types of bonds are present in an ionic compound that contains a polyatomic ion?

# Modern Chemistry Review Section 6 Answer Key

**ML Morrison**



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### **Table of Contents Modern Chemistry Review Section 6 Answer Key**

1. Understanding the eBook Modern Chemistry Review Section 6 Answer Key
  - The Rise of Digital Reading Modern Chemistry Review Section 6 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Review Section 6 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Review Section 6 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Review Section 6 Answer Key
  - Personalized Recommendations
  - Modern Chemistry Review Section 6 Answer Key User Reviews and Ratings
  - Modern Chemistry Review Section 6 Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Review Section 6 Answer Key Free and Paid eBooks

- Modern Chemistry Review Section 6 Answer Key Public Domain eBooks
  - Modern Chemistry Review Section 6 Answer Key eBook Subscription Services
  - Modern Chemistry Review Section 6 Answer Key Budget-Friendly Options
6. Navigating Modern Chemistry Review Section 6 Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Modern Chemistry Review Section 6 Answer Key Compatibility with Devices
  - Modern Chemistry Review Section 6 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Modern Chemistry Review Section 6 Answer Key
  - Highlighting and Note-Taking Modern Chemistry Review Section 6 Answer Key
  - Interactive Elements Modern Chemistry Review Section 6 Answer Key
8. Staying Engaged with Modern Chemistry Review Section 6 Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Review Section 6 Answer Key
9. Balancing eBooks and Physical Books Modern Chemistry Review Section 6 Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Review Section 6 Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Review Section 6 Answer Key
- Setting Reading Goals Modern Chemistry Review Section 6 Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Review Section 6 Answer Key
- Fact-Checking eBook Content of Modern Chemistry Review Section 6 Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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