

## CHAPTER 6 REVIEW

***The Periodic Law*****SECTION 3****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ When an electron is added to a neutral atom, energy is
  - (a) always absorbed.
  - (b) always released.
  - (c) either absorbed or released.
  - (d) neither absorbed nor released.
2. \_\_\_\_\_ The energy required to remove an electron from a neutral atom is the atom's
  - (a) electron affinity.
  - (b) electron energy.
  - (c) electronegativity.
  - (d) neither absorbed nor released.
3. From left to right across a period on the periodic table,
  - \_\_\_\_\_ a. electron affinity values tend to become more (negative or positive).
  - \_\_\_\_\_ b. ionization energy values tend to (increase or decrease).
  - \_\_\_\_\_ c. atomic radii tend to become (larger or smaller).
4. \_\_\_\_\_ a. Name the halogen with the least-negative electron affinity.  
\_\_\_\_\_ b. Name the alkali metal with the highest ionization energy.  
\_\_\_\_\_ c. Name the element in Period 3 with the smallest atomic radius.  
\_\_\_\_\_ d. Name the Group 14 element with the largest electronegativity.
5. Write the electron configuration of the following:
  - a. Na  
\_\_\_\_\_
  - b. Na<sup>+</sup>  
\_\_\_\_\_
  - c. O  
\_\_\_\_\_
  - d. O<sup>2-</sup>  
\_\_\_\_\_
  - e. Co<sup>2+</sup>  
\_\_\_\_\_

# Modern Chemistry Review Answer Sheet

**L Manion**



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