

1. List four metals that will not replace hydrogen in an acid.

Cu, Ag, Au, Pt

2. Consider the metals iron and silver, both listed in Table 3 on page 286 of the text? Which one readily forms an oxide in nature, and which one does not?

Fe forms an oxide in nature, and Ag does not, because it is much less active

3. In each of the following pairs, identify the more active element.

- a. F_2 and I_2 : **F_2**
- b. Mn and K: **K**
- c. Cu and H: **H**

4. Use the information in Table 3 on page 286 of the text to predict whether each of following reactions will occur. For each reaction that will occur, complete the chemical equation. By writing in the products formed and balancing the final equation.

- a. $Al(s) + CH_3COOH(aq) \xrightarrow{50^\circ C}$



- b. $Al(s) + H_2O(l) \xrightarrow{50^\circ C}$

No reaction

- c. $Cr(s) + CdCl_2(aq) \xrightarrow{\hspace{1cm}}$



- d. $Br_2(l) + KCl(aq) \xrightarrow{\hspace{1cm}}$

No reaction

5. Very active metals will react with water to release hydrogen gas and form hydroxides.

- a. Complete, and then balance, the equation for the reaction of $Ca(s)$ with water.



- b. The reaction of rubidium, Rb, with water is faster and more violent than the reaction of Na with water. Use the atomic structure and radius of each metal to account for this difference.

Both are alkali metals and readily form a stable $1+$ ion by ejecting $s1$ electron. Rb has a larger radius than Na and holds its electron less tightly, making it more reactive.

6. Gold, Au, is often used in jewelry. How does the relative activity of Au relate to its use in jewelry.

Gold has a low reactivity and therefore does not corrode over time.

7. Explain how to use an activity series to predict the outcome of a single-displacement reaction.

Modern Chemistry Chapter 8 Section 3 Review Answers

Kula C. Misra



Modern Chemistry Chapter 8 Section 3 Review Answers:

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