Answer Key.

Calculating Specific Heat Worksheet

Name:		Det.
		100

 $Q = mc\Delta T$, where Q = heat energy, m = mass, and $\Delta T = change in temp. Remember, <math>\Delta T = (T_{treat} - T_{initial})$. Show all work and proper units.



 A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

 How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

 100.0 g of 4.0°C water is heated until its temperature is 57°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

Mathworksheetland Answer Key

Clifford Lane



The book delves into Mathworksheetland Answer Key. Mathworksheetland Answer Key is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Mathworksheetland Answer Key, encompassing both the fundamentals and more intricate discussions.

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