

which can be determined by turning the equipment off and on, the user is encouraged to try to correct the interference by one or more of the following measures:

- Reorient or relocate the receiving antenna.
- Increase the separation between the equipment and receiver.
- Connect the equipment to an outlet on a circuit different from that to which the receiver is connected.
- Consult the dealer or an experienced radio/TV technician for help.

Important: Changes or modifications to this product not authorized by Apple could void the electromagnetic compatibility (EMC) and wireless compliance and negate your authority to operate the product. This product has demonstrated EMC compliance under conditions that included the use of compliant peripheral devices and shielded cables between system components. It is important that you use compliant peripheral devices and shielded cables between system components to reduce the possibility of causing interference to radios, televisions, and other electronic devices. This product was tested for EMC compliance under conditions that included the use of Apple peripheral devices.

Responsible party (contact for FCC matters only):

Apple Inc., Corporate Compliance
1 Infinite Loop, MS 91-1EMC
Cupertino, CA 95014

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Bluetooth Industry Canada Statement

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Fruity Loops 10 User Manual

Christian Drosten



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_2(\text{g}) \rightarrow 2 \text{ H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ/mole}$

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 $\text{C}_2\text{H}_4(\text{g}) + 2 \text{ NO}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + \text{N}_2\text{O}_5(\text{g})$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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