STD XII - CHEMISTRY

Time : 3 hrs.] OCTOBER - 2014 [Marks : 70

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question : [7]

(i) The Arrhenius equation is _____.

(a) $k = A\,e^{\frac{RT}{E_a}}$ (b) $A = k\,e^{\frac{-E_a}{RT}}$

(c) $k = A\,e^{\frac{-RT}{E_a}}$ (d) $k = A\,e^{\frac{-E_a}{RT}}$

(ii) If the enthalpy of vaporisation of water at 100oC is 186.5 J . mol–1, the entropy of vaporization will be _____.

(a) 4.0 J . K^{-1}. mol^{-1} (b) 3.0 J . K^{-1}. mol^{-1}

(c) 1.5 J . K^{-1}. mol^{-1} (d) 0.5 J . K^{-1}. mol^{-1}

(iii) The atomicity of sulphur in orthorhombic sulphur is _____.

(a) 8 (b) 6 (c) 4 (d) 2

(iv) The major binding force in diamond is _____.

(a) Covalent bond (b) Ionic bond

(c) Metallic bond (d) Co-ordinate covalent bond

(v) The boiling point of water at high altitude is low, because _____.

(a) the temperature is low. (b) the atmospheric pressure is low.

(c) the temperature is high. (d) the atmospheric pressure is high.

(vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is _____.

(a) 90 $mhos.cm^2$. (b) 110 $mhos.cm^2.mol^{-1}$

(c) 200 $mhos.cm^2.mol^{-1}$ (d) 400 $mhos.cm^2.mol^{-1}$

(vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?

(a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination

Q. 2. Answer any THREE of the following : [9]

(i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –

(a) Magnetic property (b) Oxidation state (c) Hydrides

(ii) What is the value of $\Delta S_{surr.}$ for the following reaction at 298 K –

$$6CO_{2(g)} + 6H_2O_{(l)} \longrightarrow C_6H_{12}O_{6(s)} + 6O_{2(g)}$$

Given that : $\Delta G^\circ = 2879$ kJ mol^{-1}, $\Delta S = -210$ JK^{-1} mol^{-1}.

(iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.

(iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol^{-1}) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46oC. Calculate the molar mass of the unknown compound.

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